Use Hess' Law(s) for the following reaction to: (a) determine whether exothermic/endothermic, (b) find the change in entropy and (c) determine whether spontaneous or not. Check your answer to (c) using your results from

$$\Delta H_{\text{reaction}}^{\theta} = \sum \Delta H_{f(\text{products})}^{\theta} - \sum \Delta H_{f(\text{reactants})}^{\theta}$$

$$2O_{2(g)} + CH_{4(g)} \rightarrow 2H_2O_{(g)} + CO_{2(g)}$$

 $O = -74.8 = -241.8 = -393.5$
 $205.1 = 186.3 = 188.8 = 213.7 = 0.000$

$$\Delta S^{\theta}_{\text{reaction}} = \sum S^{\theta}_{\text{(products)}} - \sum S^{\theta}_{\text{(reactants)}}$$

$$\Delta S^{\theta}_{\text{reaction}} = 2 \cdot 188, 8 + 213, 7 - (2 \cdot 205, 1 + 186, 3) = -5, 2 \frac{3}{\text{reaction}}$$

$$\Delta G^{\theta}_{\text{reaction}} = \sum \Delta G^{\theta}_{f(\text{products})} - \sum \Delta G^{\theta}_{f(\text{reactants})}$$

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