Chem 115: Answers to Chapter 6_Aqueous Ionic Equilibrium Practice Questions

Name of Instructor: Pearson W. K. Ahiahonu

Practice Question 1

Calculate the pH of the buffer that results from mixing 60.0 mL of 0.250 M $HCHO_2$ and 15.0 mL of 0.500 M $NaCHO_2(K_a)$ for $HCHO_2 = 1.8 \times 10^{-4}$)

HCHO2 [A] + H20 (1)
$$\rightleftharpoons$$
 H30 [A] + CHO2 (A)

Total volume of buffer = 60.0 ml + 15.0 ml
= 75.0 ml = 0.075 L

[HCHO2] = $\frac{0.060L}{0.075L} \times 0.250M = 0.200M$

[CHO2] = $\frac{0.015L}{0.075L} \times 0.500M = 0.100M$
 $\Rightarrow pH = pKa + log [CHO2]$
[HCHO2]
= $-log (1.8 \times 10^{-4}) + log (\frac{0.100}{0.200})$
= $-(0.25527 - 4) + (-0.30103)$
= $3.74 + 7275 - 0.30103$

Aqueous Ion Equilibrium Practice

Catherine N. Mulligan, Masaharu Fukue, Raymond N. Yong

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