

Stoichiometry Lab: Vinegar and Baking Soda

Do Not Write On This Sheet

Purpose: To predict the amount of Carbon Dioxide gas that should be produced in a chemical reaction; then calculate the % yield.



Materials: Baking Soda (NaHCO_3), Vinegar (CH_3COOH), and 2 plastic cups, scale.

Procedure: Find and record the mass of cup A. With cup A still on the scale, add approximately 10.0 g of baking soda to the cup. **The mass does not have to be exact, as long as you carefully record your results.**

Place cup B on the scale, weigh and record approximately 50.0 g of vinegar. **The mass does not have to be exact, as long as you carefully record your results.**

Slowly add vinegar to cup A until the reaction has stopped. **DO NOT** add all of the vinegar, just enough to complete the reaction. Reweigh and record both cup A and B. Calculate the mass of CO_2 that escaped.

Data:

- a. Mass of Cup A _____ g
- b. Mass of Cup A and baking soda _____ g
- c. Calculate mass of baking soda (b - a) _____ g
- d. Mass of Cup B with vinegar _____ g
- e. Mass of Cup B after reaction _____ g
- f. Calculate mass of vinegar poured into Cup A (d - e) _____ g
- g. Mass of Cup A after reaction _____ g
- h. Calculate mass of product after reaction (g - a) _____ g
- i. Calculate baking soda + vinegar (c + f) _____ g
- j. Calculate mass of CO_2 lost (i - h) _____ g

Analysis Of Baking Soda Lab Answers

**Frederick A. Bettelheim, Joseph Marvin
Landesberg**



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Table of Contents Analysis Of Baking Soda Lab Answers

1. Understanding the eBook Analysis Of Baking Soda Lab Answers
 - The Rise of Digital Reading Analysis Of Baking Soda Lab Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Analysis Of Baking Soda Lab Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Analysis Of Baking Soda Lab Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Analysis Of Baking Soda Lab Answers
 - Personalized Recommendations
 - Analysis Of Baking Soda Lab Answers User Reviews and Ratings
 - Analysis Of Baking Soda Lab Answers and Bestseller Lists
5. Accessing Analysis Of Baking Soda Lab Answers Free and Paid eBooks
 - Analysis Of Baking Soda Lab Answers Public Domain eBooks
 - Analysis Of Baking Soda Lab Answers eBook Subscription Services
 - Analysis Of Baking Soda Lab Answers Budget-Friendly Options
6. Navigating Analysis Of Baking Soda Lab Answers eBook Formats

- ePub, PDF, MOBI, and More
- Analysis Of Baking Soda Lab Answers Compatibility with Devices
- Analysis Of Baking Soda Lab Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Analysis Of Baking Soda Lab Answers
 - Highlighting and Note-Taking Analysis Of Baking Soda Lab Answers
 - Interactive Elements Analysis Of Baking Soda Lab Answers
- 8. Staying Engaged with Analysis Of Baking Soda Lab Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Analysis Of Baking Soda Lab Answers
- 9. Balancing eBooks and Physical Books Analysis Of Baking Soda Lab Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Analysis Of Baking Soda Lab Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Analysis Of Baking Soda Lab Answers
 - Setting Reading Goals Analysis Of Baking Soda Lab Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Analysis Of Baking Soda Lab Answers
 - Fact-Checking eBook Content of Analysis Of Baking Soda Lab Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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