Vitamin C Analysis (Redox Titration)

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PURPOSE: The goal of this experiment is to determine the ascorbic acid content in juice and a vitamin C tablet using redox titration with 2,6-dichloroindophenol.

Required Reading & Videos

- Ref. (I), section 4.4 (Oxidation-Reduction (Redox) Reactions), solve sample problem 4.18
- Very important: Read sections G and H of Easter Laboratory Skills and Economics review section F
- Video: Making a Solution Using a Volumetric Flask.
- Video: Using an Automatic Pipet

Introduction.

Ascorbic acid (AA), commonly known as vitamin C, is an organic compound of carbon, hydrogen and oxygen that occurs naturally in fruits and vegetables. It is an essential nutrient in the diet but is readily assidized or destroyed when emosed to heat and oxygen during food spocessing. packaging and storage. Vitamin C deficiency causes scurvy a wedespread sailor's disease until the early 19th century. with serious symptoms including bleeding of the mucous membranes, anemia and eventually death. In 1747, James Lind, a naval surgeon, conducted an experiment with sick sailors. He discovered that citrus fruits cured scurvy. In 1928, vitamin C was isolated for the first time, and five years later its chemical structure was determined. Vitamin C was named accorbic acid to indicate its anti-corbutic (antiscurvy) properties. Soon after its discovery, an industrial process was developed to produce ascerbic acid from glucose. Today it is the most produced vitamin in the world exceeding 100,000 tons per year (2).

Thrimetric Analysis of Vitamin C

AA is not only used as a dietary supplement but also as a food preservative. It acts as an anticoodant — a substance that inhibits condation and free radical damage to cells. When AA is involved in redoor reactions, it plays the role of a reducing agent by donating two electrons.

$$C_6H_6O_6 \rightarrow C_6H_6O_6 + 2H^* + 2e^-$$

 AA debydco- AA

In this reaction, AA is condined to dehydroascorbic and. The content of AA in samples can be determined by titrating the samples with 2,6-dichloroindophenol (DCP). DCP is also an organic compound and acts as an oxidizing reagent.

$$C_{13}H_3O_3NCl_3 + 2H^* + 2\pi \rightarrow C_{13}H_3O_3NCl_3$$

DCP letaco dyre

In this reaction, DCP is reduced to leuco dye. Combining these two chemical equations, we obtain the equation for the reaction of AA with DCP:

$$C_6H_6O_6 + C_DH_7O_3NCI_3 \rightarrow C_6H_6O_6 + C_DH_9O_3NCI_3$$
AA DCP debroko-AA leuco dre

DCP is a blue dye, but it is pink under acidic conditions. In a titration with blue DCP solution, the dye initially turns pink when it enters the acidic solution, but then the color is immediately removed because it is reduced to the colorless leuco dye. Once the titration is complete, a fraction of a drop of excess DCP solution imparts a pink color to the solution. The error introduced by the excess DCP is negligible. DCP is very reactive and decomposes over time. Therefore, DCP solutions must be freshly prepared and standardized just prior to use. Standardization is the titration of a titrant with a weighed quantity of a primary standard. With the known amount of the primary standard used, the concentration of the titrant can be calculated. In this experiment, you will standardize the DCP solution with reagent-grade AA.

The titration is performed in the presence of metaphosphoric acid and acetic acid (MPA-HOAc solution) to stabilize AA. Keep in mind that this experiment is quantitative. Work accurately and transfer solids and liquids quantitatively. All students must demonstrate their ability to use an automatic pipet prior to the experiment.

Apparatus:

125-ml. Erlenmeyer flasks, buret, automatic pipet & tip, 50-ml. volumetric flask, 100-ml. volumetric flask, funnel, mortar and pestle, and graduated cylinders.

Chemicals

- MPA-HOAc solution (contains 3% metaphosphoric acid and 8.4% acetic acid): Toxic and corrosive. Causes severe skin burns and eye damage.
- A recerbic acid: not hazardous.
- 2,6-Dechloroundophenol solution (the molarity of DCP will be determined in Part I). May cause eye, skin, and respiratory tract irritation. Spill on skin Immediately flush skin with excess water for 15 minutes while removing contaminated clothing. This solution also contains 0.004 M NaHCOs.

Experimental Procedure

Fart I. Standardizing the DCP Solution

- Using a 50-mL volumetric flask, prepare an aqueous solution that is approximately 0.005 M in AA (see video and required reading for instructions). Include the procedure and calculations in your notebook <u>prior to lab</u>. Record the mass of AA accurately and recalculate the molarity. Pour the solution into a beaker.
- 2. The automatic pipet (AP) at your station should only be used to transfer AA solution and Vitamin C tablet solution (Part III). Rings the tip between different solutions. The AP is a delicate and expensive instrument. Handle it carefully and follow the instructions! Transfer 2.50 mL of your AA stock solution to a 125-mL. Erlenmeyer flask using the AP at your station.

Analysis For Vitamin C Pre Lab Answers

Brigitte Niedzwiecki, Julie Pepper

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Safety-Scale Lab Exp Biochem 2e Spencer L. Seager, Michael R. Slabaugh, 1994-05

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