

BASIC CHEMISTRY – Part 2

Name: Answer Key

Date: _____

Pd: _____

PART 1: MIXTURES:

Fill in the blanks from the information provided on the presentation.

- A **mixture** is two or more substances **blended** together (**not bonded**).
- Each substance in a mixture keeps all of its **chemical** and **physical** properties.
- Mixtures, unlike compounds, can be **separated** into their individual substances by **physical** means.
- Example: **salt** and **water** can be separated by **evaporation**.
- A **solution** is a mixture of **two** or more substances, one **dissolved** in another.
- Two parts of a solution are the **solute** and the **solvent**.
- **Solute** – substance being **dissolved**
Examples: **sugar, salt**
- **Solvent** – substance doing the **dissolving**
Example: **water**
- A solution becomes **saturated** when the solvent **holds** all the solute that it can. It can **no longer dissolve** solute.
- **Suspensions** are a mixture of particles **scattered** throughout another. The particles **do not** dissolve.
Examples: **clouds, dust, fog, smoke, mud, blood**

Reflect:

1. Give another example of solutions including the solute and solvent.
Vinegar and Calcium Carbonate, etc...
2. Can solutions be separated by chemical means?
Perhaps, but will still involve physical separation as well

PART 2: COMPOUNDS

- **Compounds** are **two** or **more** elements **chemically** joined by **bonds**.
- They are represented by chemical **formulas**.
- **Properties** of compounds are usually **different** than the atoms.
Example: Hydrogen(**gas**) + Oxygen(**gas**) = Water (**liquid**)
Example: Iron + Oxygen = **Rust**
- Compounds can be **separated** into their elements only if the chemical **bonds** are **broken**. This can be done by **chemical** means (chemical **reaction**)
- A **molecule** is the **smallest** part of a compound—like water—that has **all** the **properties** of the compound.

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