

balance nuclear reaction equations

aluminum-26 undergoes positron emission
potassium-40 undergoes beta decay (electron)
uranium-238 undergoes alpha decay



electron configurations

Write the electron configuration for C.

Write the electron configuration for Cr.

Write the shorthand notation for Bi.



NOMENCLATURE [naming compounds, writing formulas]

Write the formula for each of the following compounds:

- a. Sulfur hexafluoride SF_6
- b. Lithium nitride Li_3N
- c. Chromium (III) carbonate $\text{Cr}_2(\text{CO}_3)_3$
- d. Tin (II) chloride SnCl_2
- e. Ammonium acetate $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$)
- f. Mercury (I) chloride Hg_2Cl_2
- g. Potassium bromate KBrO_3
- h. Hydrosulfuric acid $\text{H}_2\text{S(aq)}$
- i. Chloric acid HClO_3
- j. Sulfurous acid H_2SO_3

Name each of the following compounds.

- a. CuSO_4 Copper (II) sulfate
- b. AlF_3 aluminum fluoride
- c. HI hydrogen iodide
- d. NO nitrogen monoxide
- e. $\text{H}_2\text{Se(aq)}$ hydroselenic acid
- f. HNO_3 nitric acid
- g. HNO_2 nitrous acid
- h. NaHSO_4 sodium hydrogen sulfate

Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?

$$40 \xrightarrow{1} 20 \xrightarrow{2} 10 \xrightarrow{3} 5 \xrightarrow{4} 2.5 \xrightarrow{5} 1.25$$

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

$$3 \text{ half lives} \quad 100 \xrightarrow{1} 50 \xrightarrow{2} 25 \xrightarrow{3} 12.5 \text{ mg}$$

Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$$\begin{array}{rcl} 63 \times .692 & = & 43.596 \rightarrow 44 \\ 65 \times .308 & = & 20.02 \rightarrow 20. \\ \hline & + & \\ & 63.62 & \rightarrow 64 \end{array}$$

Answers For Pg 353 For Modern Chemistry

N Noddings



Answers For Pg 353 For Modern Chemistry:

Handbook of Modern Chemistry Inorganic and Organic for the Use of Students Charles Meymott Tidy, 1887
Scientific American, 1856 **Attenuation of Low Frequency Sound in the Sea**, 1980 Modern Chemical Technology and Emission Control M.B. Hocking, 2012-12-06 This text of applied chemistry considers the interface between chemistry and chemical engineering using examples of some of the important processes in industries. Integrated with this is detailed consideration of measures which may be taken for avoidance or control of potential emissions. This new emphasis in applied chemistry has been developed through eight years of experience gained from working in industry in research development and environmental control fields plus twelve years of teaching here using this approach. It is aimed primarily towards science and engineering students as well as to environmentalists and practising professionals with responsibilities or an interest in this interface. By providing the appropriate process information back to back with emissions and control data, the potential for process fine tuning is improved for both raw material efficiency and emission control objectives. This approach also emphasizes integral process changes rather than add-on units for emission control. Add-on units have their place when rapid action on an urgent emission problem is required or when control simply is not feasible by process integral changes alone. Obviously fundamental process changes for emission containment are best conceived at the design stage. However, at whatever stage process modifications are installed, this approach to control should appeal to the industrialist in particular in that something more substantial than decreased emissions may be gained.

Fundamentals of Chemistry: A Modern Introduction Frank Brescia, 2012-12-02 Fundamentals of Chemistry Third Edition introduces the reader to the fundamentals of chemistry including the properties of gases, atomic and molecular weights and the first and second laws of thermodynamics. Chemical equations and chemical arithmetic are also discussed along with the structure of atoms, chemical periodicity, types of chemical bonds and condensed states of matter. This book is comprised of 26 chapters and begins with a historical overview of chemistry and some terms which are part of the language of chemists. Separation and purification are covered in the first chapter while the following chapters focus on atomic and molecular weights, stoichiometry, the structure of atoms and types of chemical bonds. The molecular orbital (MO) theory of bonding, galvanic cells and chemical thermodynamics are considered next. Separate chapters are devoted to MO theory of covalent and metallic bonding, orbital hybridization, intermolecular forces, acids and bases, ionic equilibrium, calculations and polymers and biochemicals. This monograph is intended for chemistry students.

Test it Yourself! Lawrence F. Tuleen, Willard L. Muehl, George Sherman Porter, 1941 **Chemistry** Frank Brescia, 1974 **Chemistry To-day** Organisation for Economic Co-operation and Development, 1963 *The Educational Times, and Journal of the College of Preceptors*, 1906 **Cambridge University Examination Papers**, 1898

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