#### Stoichiometry Worksheet #2 (mole-mass, mass-mole problems)

- N<sub>2</sub> + 2O<sub>2</sub> → N<sub>2</sub>O<sub>4</sub>
  - a. If 15.0g of N<sub>2</sub>O<sub>4</sub> was produced, how many moles of O<sub>2</sub> were required?

15.0g N<sub>2</sub>O<sub>4</sub> 1 mol N<sub>2</sub>O<sub>4</sub> 2 mol O<sub>2</sub> = 0.326 mol O<sub>2</sub> 92.0g N<sub>2</sub>O<sub>4</sub> 1 mol N<sub>2</sub>O<sub>4</sub> = 0.326 mol O<sub>2</sub>

b. If 4.0x10<sup>-3</sup> moles of oxygen reacted, how many grams of N<sub>2</sub> were needed?

- K<sub>3</sub>PO<sub>4</sub> + AJ(NO<sub>3</sub>)<sub>3</sub> → 3KNO<sub>3</sub> + AJPO<sub>4</sub>
  - a. What is the mass of potassium nitrate that is produced when 2.04 moles of potassium phosphate react?

2.04 mol K<sub>3</sub>PO<sub>4</sub> 3 mol KNO<sub>3</sub> 101.1 g KNO<sub>3</sub> = 619g KNO<sub>3</sub> 1 mol K<sub>3</sub>PO<sub>4</sub> 1 mol KNO<sub>3</sub>

> b. If 5.80g of aluminum phosphate are formed, how many moles of aluminum nitrate reacted?

5.80g AIPO<sub>4</sub> 1 mol AIPO<sub>6</sub> 1 mol AI(NO<sub>1</sub>)<sub>1</sub> = 4.75×10<sup>-2</sup> mol AI(NO<sub>1</sub>)<sub>1</sub> 122.0g AIPO<sub>6</sub> 1 mol AIPO<sub>6</sub>

- CaC<sub>2</sub> + 2H<sub>2</sub>O → Ca(OH)<sub>2</sub> + C<sub>2</sub>H<sub>2</sub>
  - a. If you have 5.50 mol of CaC<sub>2</sub>, how much C<sub>2</sub>H<sub>2</sub> do you get?

5.50 mol CaC<sub>2</sub> 1 mol C<sub>2</sub>H<sub>2</sub> 26.0g C<sub>2</sub>H<sub>2</sub> = 143g C<sub>2</sub>H<sub>2</sub> 1 mol CaC<sub>2</sub> 1 mol C<sub>2</sub>H<sub>2</sub>

b. How many moles of water are needed when 65.0g of CaC<sub>2</sub> have reacted?

65.0g CaC<sub>2</sub> 1 mol CaC<sub>2</sub> 2 mol H<sub>2</sub>O = 2.03 mol H<sub>2</sub>O 64.1g CaC<sub>2</sub> 1 mol CaC<sub>2</sub>

4. In photosynthesis, water reacts with carbon dioxide to give oxygen and glucose (C<sub>6</sub>H<sub>12</sub>O<sub>6</sub>). Write and balance the chemical equation. How many moles of CO<sub>2</sub> are required to make 120.0g of glucose?

6H2O + 6CO2 -- C4H22O4 + 6O2

120.0g C<sub>6</sub>H<sub>12</sub>O<sub>8</sub> 1 mol C<sub>6</sub>H<sub>12</sub>O<sub>8</sub> 6 mol CO<sub>2</sub> = 4.000 mol CO<sub>2</sub> 180.0 g C<sub>6</sub>H<sub>12</sub>O<sub>8</sub> 1 mol C<sub>6</sub>H<sub>12</sub>O<sub>6</sub>

# **Answers For The Chemistry Lab Of Stoichiometry**

**Jiyuan Zhang** 

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