

Lab 10: Calculations with a Chemical Structure

Objectives:

1. To derive an empirical formula from a chemical structure and various percentages forming (molecular) compounds.
2. To calculate the number of moles of each of the components in the mixture.
3. To determine the molar mass of the mixture.
4. To determine the theoretical amount of product that would be formed.
5. To compare the theoretical amount to the actual amount of product formed and calculate the percent yield.

Reaction equation:

Reaction of sodium chloride solution: 25.0 mL

Reaction of calcium chloride solution: 25.0 mL

Reaction of 10% NaOH: 5.00 g

Reaction of 10% NaOH: 5.00 g

Reaction of 10% NaOH: 5.00 g

Reaction of 10% NaOH: 5.00 g

Procedure:

Reaction and Calculations:

1. Reaction:

$$\text{mole of sodium chloride} = 25.0 \text{ mL} \times \frac{1.0 \text{ g/mL}}{58.44 \text{ g/mol}} = 0.428 \text{ mol}$$

$$\text{mole of calcium chloride} = 25.0 \text{ mL} \times \frac{1.0 \text{ g/mL}}{110.98 \text{ g/mol}} = 0.225 \text{ mol}$$

$$\text{mole of sodium hydroxide} = 5.00 \text{ g} \times \frac{1 \text{ mol}}{40.00 \text{ g/mol}} = 0.125 \text{ mol}$$

$$\text{mole of product} = 0.125 \text{ mol} \times \frac{1 \text{ mol}}{1 \text{ mol}} = 0.125 \text{ mol}$$

2. Reaction:

3. Reaction:

$$\text{mole of } \text{CaCl}_2 = 0.225 \text{ mol} \times \frac{1 \text{ mol}}{1 \text{ mol}} = 0.225 \text{ mol}$$

$$\text{mole of } \text{CaCl}_2 = 0.225 \text{ mol} \times \frac{1 \text{ mol}}{1 \text{ mol}} = 0.225 \text{ mol}$$

$$\text{mole of } \text{CaCl}_2 = 0.225 \text{ mol} \times \frac{1 \text{ mol}}{1 \text{ mol}} = 0.225 \text{ mol}$$

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