

201secondary Solutions Odyssey Literature Guide

Answer Key

Joacim Rocklöv



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denoted by the letter z which is defined as the number of units of positive charge protons in the nucleus for example if an
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these substances exchange atoms with each other to form new products through chemical reactions key

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web the atomic number tells you the number of of an n one atom of an element it also tells you the number of in a neutral atom of that element the atomic number gives the identity of an element as well as ts location on the periodic table no two different elements will have the give the symbol and number of protons in one atom of atomic number

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web overall the experiment succeeded in showing that temperature and pressure for an ideal gas at constant volume and mass follow the relation of the ideal gas equation differences existed in the experimental graph of temperature versus and pressure and the theoretical curve of temperature versus pressure

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web we then changed the volume between 5ml and 20ml and recorded the pressure at each this part of the lab showed that volume and pressure are inversely proportionate as the pressure would decrease if the volume was increased part b of the experiment looked at the relationship between temperature and pressure

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web procedure 1 the relationship between pressure and volume set the plunger of the syringe to 1 10 the capacity of the syringe the 2 0 ml mark if using a 20 ml syringe the volume is read from the edge of the bottom rib of the plunger attach the syringe to the connector at the end of sensor

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web he concluded that at a constant pressure the volume of the gas would expand at a particular rate for each degree of temperature rise that being 1 267 per degree in 1874 victor regnault obtained better experimental results showing this number to be 1 273 and concluded that the pressure would approach zero at 1 273 degrees c

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october 8 2014 summary the purpose of this experiment is to be able to understand the relationship between the pressure
volume and temperature of a gas understanding this relationship is important as it allows

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web the volume of 1 00 mol 1 00 mol of any gas at stp standard temperature 273 15 k and pressure 1 atm is measured to be
22 414 l 22 414 l we can substitute 101 325 kpa 101 325 kpa for pressure 22 414 l 22 414 l for volume and 273 15 k 273 15 k
for temperature into the ideal gas equation and solve for r r

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web chemists sometimes make comparisons against a standard temperature and pressure stp for reporting properties of
gases 273 15 k and 1 atm 101 325 kpa 1 at stp one mole of an ideal gas has a volume of about 22 4 l this is referred to as the
standard molar volume figure 9 18

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web this lab explores the relationships between pressure volume and temperature in an ideal gas the ideal gas law $pV = nRT$ is
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web figure page index 10 since the number of moles in a given volume of gas varies with pressure and temperature changes
chemists use standard temperature and pressure 273 15 k and 1 atm or 101 325 kpa to report properties of gases

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introduction the ideal gas law is a commonly used formula in chemistry it relates pressure temperature and volume this
experiment seeks to prove those relationships by measuring the pressure of gas at various temperatures

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properties of fluids under a wide range of pressure and temperature

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web volume pressure kpa linear pressure kpa volume ml pressure kpa in part b we observed the relationship between temperature and pressure the pressure of the 125 ml flask was at 101.36 kpa when the temperature was at 23.2 °C

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web pressure increases even though there were errors in the experiment the goal of the experiment was met the goal of declaring what type of relationship it is was clear using the data pressure kpa temperature °C temperature K constant K p t 95.82 kpa 16.0 °C 289 K 0.33 97.90 kpa 19.0 °C 292 K 0.34 98.97 kpa 99.22 kpa 101.4 kpa 19.1 °C

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web the properties of water at constant volume can be represented as a function of pressure and temperature as shown in the diagram below figure 1 the pressure temperature relationship at constant volume the saturation point of water is the condition at which a phase change occurs from liquid to vapour or vapour to liquid

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web figure pageindex 5 the empirically determined relationships among pressure volume temperature and amount of a gas the thermometer and pressure gauge indicate the temperature and the pressure qualitatively the level in the flask indicates the volume and the number of particles in each flask indicates relative amounts