

Name Taylor Hoke Section _____

Experiment 1

Advance Study Assignment: The Densities of Solids and Liquids

The advance study assignments in this laboratory manual are designed to assist you in making the calculations required in the experiment you will be doing. We do this by furnishing you with sample data and showing in some detail how these data can be used to obtain the desired results. In the advance study assignments we will often include the guiding principles as well as the specific relationships to be employed. If you work through the steps in each calculation by yourself, you should have no difficulty when you are called upon to make the necessary calculations on the basis of the data you obtain in the laboratory.

1. Finding the volume of a flask.

A student obtained a clean dry glass-stoppered flask. She weighed the flask and stopper on an analytical balance and found the total mass to be 31.601 g. She then filled the flask with water and obtained a mass for the full stoppered flask of 60.735 g. From these data, and the fact that at the temperature of the laboratory the density of water was 0.9973 g/cm³, find the volume of the stoppered flask.

- a. First we need to obtain the mass of the water in the flask. This is found by recognizing that the mass of a sample is equal to the sum of the masses of its parts. For the filled stoppered flask:

Mass of filled stoppered flask = mass of empty stoppered flask + mass of water,
so mass of water = mass of filled flask - mass of empty flask

$$\text{Mass of water} = \underline{60.735\text{ g}} - \underline{31.601\text{ g}} = \underline{29.134\text{ g}}$$

Many mass and volume measurements in chemistry are made by the method used in 1a. This method is called measuring by difference, and is a very useful one.

- b. The density of a pure substance is equal to its mass divided by its volume:

$$\text{Density} = \frac{\text{mass}}{\text{volume}} \quad \text{or} \quad \text{volume} = \frac{\text{mass}}{\text{density}}$$

The volume of the flask is equal to the volume of the water it contains. Since we know the mass and density of the water, we can find its volume and that of the flask. Make the necessary calculation.

$$V = \frac{29.1349}{0.997391\text{ cm}^3}$$

$$\text{Volume of water} = \text{volume of flask} = \underline{29.21\text{ cm}^3}$$

2. Finding the density of an unknown liquid.

Having obtained the volume of the flask, the student emptied the flask, dried it, and filled it with an unknown whose density she wished to determine. The mass of the stoppered flask when completely filled with liquid was 56.796 g. Find the density of the liquid.

- a. First we need to find the mass of the liquid by measuring by difference:

$$\text{Mass of liquid} = \underline{56.796\text{ g}} - \underline{31.601\text{ g}} = \underline{29.195\text{ g}}$$

(continued on following page)

Advance Study Assignment Answers Experiment 1 Chemistry

Karin Nielsen-Saines



Advance Study Assignment Answers Experiment 1 Chemistry:

Environmental Laboratory Exercises for Instrumental Analysis and Environmental Chemistry Frank M.

Dunnivant, 2004-08-23 A comprehensive set of real world environmental laboratory experiments This complete summary of laboratory work presents a richly detailed set of classroom tested experiments along with background information safety and hazard notes a list of chemicals and solutions needed data collection sheets and blank pages for compiling results and findings This useful resource also Focuses on environmental i e dirty samples Stresses critical concepts like analysis techniques and documentation Includes water air and sediment experiments Includes an interactive software package for pollutant fate and transport modeling exercises Functions as a student portfolio of documentation abilities Offers instructors actual samples of student work for troubleshooting notes on each procedure and procedures for solutions preparation **A**

Laboratory Manual of General Chemistry for Use in Colleges William Crowell Bray, Wendell Mitchell Latimer, 1921

Chemical Investigations Nancy Konigsberg Kerner, 1986 **TEACHING OF BIOLOGICAL SCIENCES (Intended for Teaching of Life Sciences, Physics, Chemistry and General Science)** AHMAD, JASIM, 2011-11-30 Students of today especially at the school level perceive science as a collection of facts to be memorized whereas in reality it is constantly changing as new information accumulates and new techniques develop every day The objective of teaching is not restricted to imparting scientific information to students but also to help them apply these principles in their daily lives This comprehensive book written in an easy to understand language covers the entire syllabus of teaching of Biological Sciences in particular and Science Teaching in general In so doing it takes into account the needs of teacher trainees and in service teachers Organized into 20 chapters the book discusses in detail the many facets and aspects of Biology Science Teaching The text introduces modern approaches to teaching with the aim of improving student learning throughout their course It emphasizes the need for pedagogical analysis vis vis subject teaching constructive approach laboratory work Continuous and Comprehensive Evaluation CCE In addition the text highlights the difference between microteaching and simulated teaching It also shows how e learning and co curricular activities can be successfully integrated in biological sciences teaching **NEW TO THIS EDITION** Inclusion of one chapter on Concept Mapping in Biology Teaching This chapter advocates the popularized constructivist approach of teaching learning process Besides some figures tables and flow charts are also added to make the book more useful to the readers **KEY FEATURES** Analyses Constructivism versus Behaviourism Includes self explanatory model lesson plan Discusses Information and Communication Technology ICT in the context of Biology Science teaching learning Suggests how apparatus and devices can be secured and cultured and used in classroom demonstrations and student projects Primarily intended as a text for students of B Ed pursuing course on Teaching of Biological Sciences Life Sciences the book should prove equally useful for B Ed students following courses on Teaching of Physical Sciences In addition diploma students of Elementary Teacher Education ETE having a paper on Teaching of EVS General Science and M

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