

Advance Study Assignment: Identification of a Compound by Mass Relationships

1. A student attempts to identify an unknown compound by the method used in this experiment. She finds that when she heated a sample weighing 0.4862 g the mass went down appreciably, to 0.3067 g. When the product was converted to a chloride, the mass went up, to 0.3380 g.

a. Is the sample a carbonate?

Indicate Yes

NO

Give your reasoning.

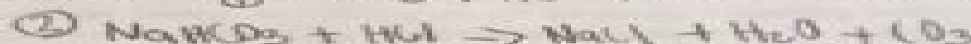
b. What are the two compounds that might be in the unknown?

Sodium bicarbonate

Potassium



c. Write the chemical equation for the overall reaction that would occur when the original compound was converted to a chloride. If the compound is a hydrogen carbonate, use the sum of Reactions 1 and 2. If the sample is a carbonate, use Reaction 2. Write the equation for a sodium salt and then for a potassium salt.



d. How many moles of the chloride salt would be produced from one mole of original compound?

1 mol of chloride salt would be formed from 1 mol of HCO_3^- ; 2 mol of Cl⁻ salt would be formed from 1 mol of CO_3^{2-} .

e. How many grams of the chloride salt would be produced from one molar mass of original compound?

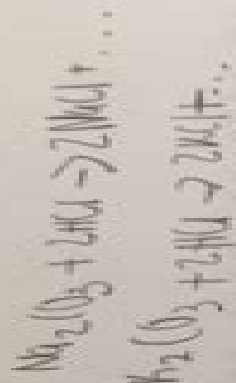
Molar masses: NaHCO_3 84.007 Na_2CO_3 105.99 NaCl 58.44
 KHCO_3 100.087 K_2CO_3 138.205 KCl 74.551
 If a sodium salt, 0.250 g original compound $\rightarrow$ 0.302 g chloride
 If a potassium salt, 0.250 g original compound $\rightarrow$ 0.278 g chloride

f. What is the theoretical value of Q, as found by Equation 3,

if she has an Na salt? 0.302 if she has a K salt? 0.278

g. What was the observed value of Q?

h. Which compound did she have as an unknown?



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