

Advance Study Assignment: Determination of a Chemical Formula

1. To find the mass of a mole of an element, one looks up the atomic mass of the element in a table of atomic masses (see Appendix III or the Periodic Table). The molar mass of an element is simply the mass in grams of that element that is numerically equal to its atomic mass. For a compound substance, the molar mass is equal to the mass in grams that is numerically equal to the sum of the atomic masses in the formula of the substance. Find the molar mass of

Cu _____ g Cl _____ g H _____ g O _____ g H₂O _____ g

2. If one can find the ratio of the number of moles of the elements in a compound to one another, one can find the formula of the compound. In a certain compound of copper and oxygen, Cu_xO_y, we find that a sample weighing 0.5424 g contains 0.4831 g Cu.

- a. How many moles of Cu are there in the sample?

$$\text{moles of Cu} = \frac{0.4831 \text{ g}}{63.5 \text{ g/mol}} = \left(\text{No. moles} = \frac{\text{mass Cu}}{\text{molar mass Cu}} \right)$$

The molar masses:
 Cu → 63.5 g/mol
 Cl → 35.5 g/mol
 H → 1 g/mol
 O → 16 g/mol

0.00760 moles

- b. How many grams of O are there in the sample? (The mass of the sample equals the mass of Cu plus the mass of O.)

$$m(O) = 0.5424 \text{ g} - 0.4831 \text{ g}$$

0.0593 g

- c. How many moles of O are there in the sample?

$$\text{moles of O} = \frac{0.0593 \text{ g}}{16 \text{ g/mol}}$$

0.00760 moles

- d. What is the mole ratio (no. moles Cu/no. moles O) in the sample?

$$\frac{0.00760 \text{ mol Cu}}{0.00760 \text{ mol O}} = \frac{\text{no. of moles Cu}}{\text{no. of moles O}} = 1 : 1$$

- e. What is the formula of the oxide? (The atom ratio equals the mole ratio, and is expressed using the smallest integers possible.)

CuO

- f. What is the molar mass of the copper oxide?

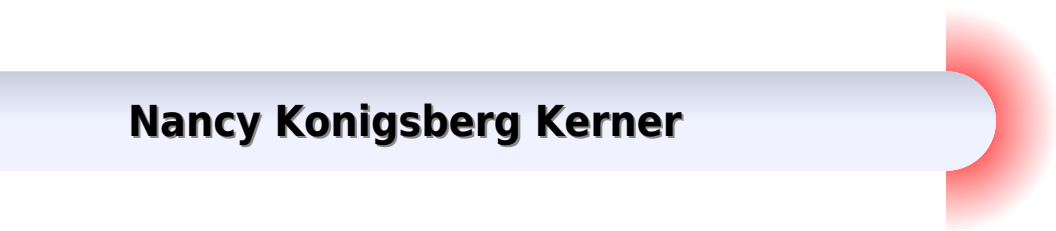
The molar mass of CuO

$$63.5 + 16 = 79.5 \text{ g/mol}$$

79.5 g

Advance Study Assignment Answers Experiment Chemistry

Nancy Konigsberg Kerner



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A Laboratory Manual of General Chemistry for Use in Colleges William Crowell Bray, Wendell Mitchell Latimer, 1921

Environmental Laboratory Exercises for Instrumental Analysis and Environmental Chemistry Frank M. Dunnivant, 2004-08-23 A comprehensive set of real world environmental laboratory experiments This complete summary of laboratory work presents a richly detailed set of classroom tested experiments along with background information safety and hazard notes a list of chemicals and solutions needed data collection sheets and blank pages for compiling results and findings This useful resource also Focuses on environmental i e dirty samples Stresses critical concepts like analysis techniques and documentation Includes water air and sediment experiments Includes an interactive software package for pollutant fate and transport modeling exercises Functions as a student portfolio of documentation abilities Offers instructors actual samples of student work for troubleshooting notes on each procedure and procedures for solutions preparation

Community and Junior College Journal , 1974

Introduction to Chemical Principles Susan A. Weiner, Peters, Edward I. Peters, 1997 This laboratory manual provides a detailed chemical overview that enables students to truly understand the function of the laboratory All experiments have been thoroughly student tested and include step by step instructions including safety and disposal methods for each

Learning and Understanding National Research Council, Division of Behavioral and Social Sciences and Education, Center for Education, Committee on Programs for Advanced Study of Mathematics and Science in American High Schools, 2002-08-06 This book takes a fresh look at programs for advanced studies for high school students in the United States with a particular focus on the Advanced Placement and the International Baccalaureate programs and asks how advanced studies can be significantly improved in general It also examines two of the core issues surrounding these programs they can have a profound impact on other components of the education system and participation in the programs has become key to admission at selective institutions of higher education By looking at what could enhance the quality of high school advanced study programs as well as what precedes and comes after these programs this report provides teachers parents curriculum developers administrators college science and mathematics faculty and the educational research community with a detailed assessment that can be used to guide change within advanced study programs

Advances in Organometallic Chemistry Anthony F. Hill, Mark J. Fink, 2011-08-09 Almost all branches of chemistry and material science now interface with organometallic chemistry the study of compounds containing carbon metal bonds This widely acclaimed serial contains authoritative reviews that address all aspects of organometallic chemistry a field that has expanded enormously since the publication of Volume 1 in 1964 Fully updated and expanded to reflect recent advances Illustrated with pertinent examples from recent literature Contributions from leading authorities and industry experts

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