Post-Lab Questions

- 1. From the experiment, which reactant is the limiting reactant? Which one is the reactant in excess? How do you determine which one is the limiting reactant?
- 2. If you weigh the watch glass then write your name on it with the wax pencil, how will that affect your percent yield?
- 3. What factors could cause the percent yield of a product to be less than 100%?
- 4. What factors could cause the percent yield of a product to be more than 100%?
- Balance the following reaction: ___C₃H₈ + ___O₂ → ___CO₂ + ___H₂O
 - a. 4.30 g of C₃H₈ reacts with 4.30 g of O₂. Which one is the limiting reagent, and which one is in excess? Explain your reasoning.
 - b. Based on the limiting reagent, what are the theoretical yields of CO₂ and H₂O?
 - c. You perform the experiment and produce 2.67 g of CO₂ and 0.79 g of H₂O. What are your percent yields for each product?

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Puja Mehta

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