

Lab 10: Calculations with a Chemical Structure

Objectives:

1. To derive an empirical formula from a chemical structure without performing stoichiometric calculations.
2. To calculate the number of moles of a solid from a weighing measurement in the lab.
3. To determine the molar mass of a solid.
4. To determine the theoretical amount of product that would be produced.
5. To compare the theoretical amount to the actual amount of substance formed and calculate the percent yield.

Reaction equation:

Reaction of sodium chloride solution: 25.0 mL

Reaction of calcium chloride solution: 25.0 mL

Reaction 10g Silver Paper: 0.00 g

Reaction 5 Silver Paper + 20g Silver: 0.00 g

Reaction of 10g Silver Paper + 20g Silver: 0.00 g

Reaction of 10g Silver Paper + 20g Silver: 0.00 g

Results:

Observations and Calculations:

1. Calculations:

$$\text{moles of sodium chloride} = 0.00 \text{ mol NaCl} \times \frac{1.0}{100.0} = 0.000000 \text{ mol NaCl}$$

$$\text{moles of calcium chloride} = 0.00 \text{ mol CaCl}_2 \times \frac{1.0}{100.0} = 0.000000 \text{ mol CaCl}_2$$

$$\text{moles of silver chloride produced} = 0.00 \text{ mol AgCl} \times \frac{1.0}{100.0} = 0.000000 \text{ mol AgCl}$$

$$\text{moles of silver chloride produced} = 0.00 \text{ mol AgCl} \times \frac{1.0}{100.0} = 0.000000 \text{ mol AgCl}$$

2. Chemical Equations: $\text{Na}_2\text{CO}_3 + \text{CaCl}_2 \rightarrow \text{CaCO}_3 + 2\text{NaCl}$

3. Silver Paper:

$$\text{Silver: } 0.00 \text{ g Ag} \times \frac{1.0}{100.0} = 0.000000 \text{ g Ag} \times \frac{1.0}{100.0} = 0.000000 \text{ g Ag}$$

$$\text{Silver: } 0.00 \text{ g Ag} \times \frac{1.0}{100.0} = 0.000000 \text{ g Ag}$$

$$\text{Silver: } 0.00 \text{ g Ag} \times \frac{1.0}{100.0} = 0.000000 \text{ g Ag} \times \frac{1.0}{100.0} = 0.000000 \text{ g Ag}$$

Answers For Lab 6b Heath Chemistry

Ruth E. McCall



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