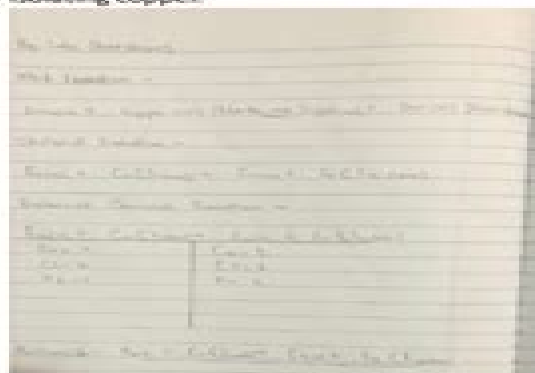


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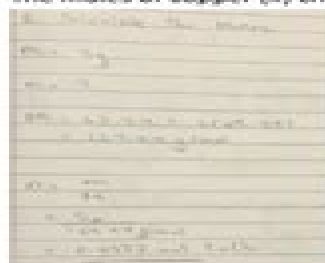
The percentage yield of a chemical reaction is determined by numerous factors: the nature of the reaction itself, the conditions under which the reaction was carried out and the nature of the reactants used. In this investigation, you will determine the limiting/excess reagents and the percentage yield of the reaction between iron and a solution of copper (II) chloride.

Prelab Questions:

- Write the balanced chemical equation for the reaction between iron and copper (II) chloride. Remember to use your activity series to check the products.
 - The balanced equation is $\text{Fe(s)} + \text{CuCl}_2\text{(aq)} \rightarrow \text{Cu(s)} + \text{FeCl}_2\text{(aq)}$. The reaction is a single displacement isolating copper.



- Calculate the moles of copper (II) chloride in the solution if you use 5 g of copper (II) chloride.
 - The moles of copper (II) chloride in the solution, if I use 5g of copper (II) chloride, is 0.037 mol CuCl_2 .



- Determine the limiting reagent if you use 1 g of iron.
 - The limiting reagent, if I use 1g of iron, is Fe(s) . It is 0.018mol because when comparing the numbers by the

Limiting Reagents and % Yield



Answers For Quick Lab Limiting Reagents Section 12 3

Patrick Vollmar



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web nov 9 2023 creating a mash the first thing you ll need for making moonshine is a mash this part of the process will depend on the flavor you want weigh and measure out all your ingredients place your mash pot on your heat source and turn it on pour in 5 gallons of water and boil it to 165 f once it reaches 165 f turn off your heat source

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web 1 make the mash the process begins by heating 5 gallons of water to 165 f when the temperature reaches this point turn off the heat and add the full amount of corn to the water it s important to stir the corn consistently for 5 minutes