Limiting Reagent & Percent Yield Practice Worksheet

- When copper (II) chloride reacts with sodium nitrate, copper (II) nitrate and sodium chloride are formed.
 - a. Write the balanced equation for the reaction given above:

+ NaNO₁ → Cu(NO₁)₂ + ChrCile CuCl₂ + 2 NaNO₃ → Cu(NO₃)₂ + 2 NaCl

 If 15 grams of copper (II) chloride react with 20 grams of sodium nitrate, how much sodium chloride can be formed?

1 mol CuCl₂ x 2 mol NaCl x 58.443 g NaCl = 13 g NaCl 15 g CuClax 134.362g CuCl: 1 mol CuCl: 1 mol NaCl

20 g NaNO₃ x 1 mol NaNO₃ x 2 mol NaCl x 58.443 g NaCl = 13.8 g NaCl 84,997g NaNOs 2 mol NaNOs 1 mol NaCl

Only 13 g can be formed

- c. What is the limiting reagent for the reaction in #2? copper (II) chloride
- d. How many grams of copper (II) nitrate is formed?

 $15g CuCl_2 \times 1 \mod CuCl_2 \times 1 \mod Cu(NO_3)_2 \times 187.467 g Cu(NO_3)_2 = 20.9 g Cu(NO_3)_2$ 134,452g CuCl₂ 1 mol CuCl₂ 1 mol Cu(NO₃)₂

e. How much of the excess reasent is left over in this reaction?

15g CuCl₂ x 1 mol CuCl₂ x 2 mol NaNO₃ x 84.997g NaNO₃ = 19 g NaNO₃ 134.452g CuCl₂ 1 mol CuCl₂ 1 mol NaNO₃

20 g NaNOs - 19 g NaNOs = 1 g NaNOs

f. If 11.3 grams of sodium chloride are formed in the reaction described in problem #2, what is Percent yield = $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100 \text{ percent}$ the percent yield of this reaction?

formed.

11.3/13.0 x 100% = 86.9%

- 2. When lead (II) nitrate reacts with sodium iodide, sodium nitrate and lead (II) iodide are
 - a. Balance the following equation:

Pb(NO₃)_{2 (and} + NaI (and → PbI_{2 (a)} + NaNO_{3 (and} Pb(NO₃)_{2 tant} + 2 NaI _{tant} → PbI_{2 tot} + 2 NaNO_{3 tant}

 If I start with 25.0 grams of lead (II) nitrate and 15.0 grams of sodium iodide, how many grams of sodium nitrate can be formed?

25.0 g Pb(NO₃)₂ x 1 mol Pb(NO₃)₂ x 2 mol NaNO₃ x 84.997g NaNO₃ = 12.8 g NaNO₃ 331.214 g Pb(NO₃)₂ 1 mol Pb(NO₃)₂ 1 mol NaNO₃

15.0 g Nal x 1 mol Nal x 2 mol NaNO₃ x 84.997g NaNO₃ = 8.51 g NaNO₃ 149,89g Nat 2 mol NaI 1 mol NaNOs

8.51 g NaNO3 can be made

Answers Limiting Reactants And Percent Yield

Mr. Rohit Manglik

Answers Limiting Reactants And Percent Yield:

Microbiology Question & Answer Purshotam Kaushik, 2008 The revised edition as per UGC model for B Sc Pass Honours and M Sc students of all Indian Universities and also useful for competitive examinations like NET GATE etc New chapters added on Human Immunodeficiency virus and AIDS Ecological Groups of Microorganisms Extremophiles Aeromicrobiology Biogeochemical Cycling and Pharmaceutical and Microbial Technology besides many illustrations The text has been made more informative. The special features include development of microbiology in the field has been provided microbiology applications the concept of microbiology bacterial nomenclature modern trends in between etc Calculating Theoretical Yield Is Not Impossible! Michael S. Leonard, 2025-02-25 Calculating theoretical and percent yield is a fundamental skill for the laboratory This book primarily targets Organic Chemistry Laboratory courses at the high school or college and university level as a supplemental resource to help students master this skill It begins with simple examples from everyday life demonstrates the importance of balancing the equation addresses the role of the mole in these computations discusses different types of liquids considers the role of significant figures and culminates with the planning of syntheses There are suggestions for further reading as well as practice problems and questions to ensure mastery Begins with examples from everyday life that enable students to understand the concepts of theoretical and percent yield before applying those concepts to the laboratory Addresses the necessity of balancing the reaction equation the centrality of the mole in these calculations and the role of significant figures in reporting the answer Explains how to approach the calculations when using neat liquids or solutions The culmination of this text is the use of the same thought processes to plan the amounts of reactants needed for syntheses of desired quantities of product All of the problems in the book include detailed solutions with accompanying text to explain the answers and ancillaries also include suggestions for further reading *Environmental* Chemistry and Green Chemistry Mr. Rohit Manglik, 2023-07-23 Addresses chemical processes affecting the environment and introduces green chemistry principles for sustainable practices CliffsStudySolver: Chemistry Charles Henrickson, 2007-05-03 The CliffsStudySolver workbooks combine 20 percent review material with 80 percent practice problems and the answers to help make your lessons stick CliffsStudySolver Chemistry is for students who want to reinforce their knowledge with a learn by doing approach Inside you ll get the practice you need to learn Chemistry with problem solving tools such as Clear concise reviews of every topic Practice problems in every chapter with explanations and solutions A diagnostic pretest to assess your current skills A full length exam that adapts to your skill level A glossary examples of calculations and equations and situational tasks can help you practice and understand chemistry This workbook also covers measurement chemical reactions and equations and matter elements compounds and mixtures Explore other aspects of the language including Formulas and ionic compounds Gases and the gas laws Atoms The mole elements and compounds Solutions and solution concentrations Chemical bonding Acids bases and buffers Practice makes perfect and whether you re

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