

Chem 115: Answers to Chapter 6_Aqueous Ionic Equilibrium Practice Questions

Name of Instructor: Pearson W. K. Ahiaonu

Practice Question 1

Calculate the pH of the buffer that results from mixing 60.0 mL of 0.250 M HCHO_2 and 15.0 mL of 0.500 M NaCHO_2 (K_a for $\text{HCHO}_2 = 1.8 \times 10^{-4}$)



$$\begin{aligned}\text{Total Volume of buffer} &= 60.0 \text{ mL} + 15.0 \text{ mL} \\ &= 75.0 \text{ mL} = 0.075 \text{ L}\end{aligned}$$

$$[\text{HCHO}_2] = \frac{0.060 \text{ L} \times 0.250 \text{ M}}{0.075 \text{ L}} = 0.200 \text{ M}$$

$$[\text{CHO}_2^-] = \frac{0.015 \text{ L} \times 0.500 \text{ M}}{0.075 \text{ L}} = 0.100 \text{ M}$$

$$\Rightarrow \text{pH} = \text{p}K_a + \log \frac{[\text{CHO}_2^-]}{[\text{HCHO}_2]}$$

$$= -\log (1.8 \times 10^{-4}) + \log \left(\frac{0.100}{0.200} \right)$$

$$= - (0.25527 - 4) + (-0.30103)$$

$$= 3.7447275 - 0.30103$$

$$= 3.44$$

Aqueous Ion Equilibrium Practice

Richard Bailey



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