# Chem 115: Answers to Chapter 6\_Aqueous Ionic Equilibrium Practice Questions

# Name of Instructor: Pearson W. K. Ahiahonu

# Practice Question 1

Calculate the pH of the buffer that results from mixing 60.0 mL of 0.250 M  $HCHO_2$  and 15.0 mL of 0.500 M  $NaCHO_2(K_a)$  for  $HCHO_2 = 1.8 \times 10^{-4}$ )

HCHO2 [A] + H20 (1) 
$$\rightleftharpoons$$
 H30 [A] + CHO2 (A)

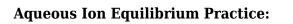
Total volume of buffer = 60.0 ml + 15.0 ml
= 75.0 ml = 0.075 L

[HCHO2] =  $\frac{0.060L}{0.075L} \times 0.250M = 0.200M$ 

[CHO2] =  $\frac{0.015L}{0.075L} \times 0.500M = 0.100M$ 
 $\Rightarrow pH = pKa + log [CHO2]$ 
[HCHO2]
=  $-log (1.8 \times 10^{-4}) + log (\frac{0.100}{0.200})$ 
=  $-(0.25527 - 4) + (-0.30103)$ 
=  $3.74 + 7275 - 0.30103$ 

# **Aqueous Ion Equilibrium Practice**

**Richard Bailey** 



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