

Experiment 2 - Gravimetric Analysis of a Chloride Salt (G)

Purpose:

Determining quantitative value of chloride in an unknown soluble salt by showing and explaining common methods in gravimetric analysis.

Theory:

- Reaction between silver ion and chloride ion to form silver chloride as an equation:



- AgCl solubility in water and the K_{sp} :

Silver chloride is very insoluble, so we can assume that the above reaction goes to completion.

$$K_{sp} = 1.6 \times 10^{-10}$$

- AgCl is precipitated with an acidic solution to reduce likelihood of acid ions interfering with results. These ions will form co-precipitates with silver in the acid, due to these co-precipitates, the results would be higher.
- To prevent the precipitate from going through the filtered paper, it is heated to coagulate it into a different texture that's easier to work with. Heating the precipitate makes it coagulate, therefore, it is harder for the precipitate to go through the filtered paper.
- Dry silver chloride decomposes into silver and chlorine in the presence of light:



- If photodecomposition takes place in air, results will be lower, since chlorine will be released in the form of gas. If photochemical decomposition occurs in the presence of excess silver ions in aqueous solution, it will result in higher results.

Analysis Of An Unknown Salt Lab Report

R Barnett



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Table of Contents Analysis Of An Unknown Salt Lab Report

1. Understanding the eBook Analysis Of An Unknown Salt Lab Report
 - The Rise of Digital Reading Analysis Of An Unknown Salt Lab Report
 - Advantages of eBooks Over Traditional Books
2. Identifying Analysis Of An Unknown Salt Lab Report
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Analysis Of An Unknown Salt Lab Report
 - User-Friendly Interface
4. Exploring eBook Recommendations from Analysis Of An Unknown Salt Lab Report
 - Personalized Recommendations
 - Analysis Of An Unknown Salt Lab Report User Reviews and Ratings
 - Analysis Of An Unknown Salt Lab Report and Bestseller Lists
5. Accessing Analysis Of An Unknown Salt Lab Report Free and Paid eBooks
 - Analysis Of An Unknown Salt Lab Report Public Domain eBooks
 - Analysis Of An Unknown Salt Lab Report eBook Subscription Services
 - Analysis Of An Unknown Salt Lab Report Budget-Friendly Options
6. Navigating Analysis Of An Unknown Salt Lab Report eBook Formats

- ePub, PDF, MOBI, and More
- Analysis Of An Unknown Salt Lab Report Compatibility with Devices
- Analysis Of An Unknown Salt Lab Report Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Analysis Of An Unknown Salt Lab Report
 - Highlighting and Note-Taking Analysis Of An Unknown Salt Lab Report
 - Interactive Elements Analysis Of An Unknown Salt Lab Report
- 8. Staying Engaged with Analysis Of An Unknown Salt Lab Report
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Analysis Of An Unknown Salt Lab Report
- 9. Balancing eBooks and Physical Books Analysis Of An Unknown Salt Lab Report
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Analysis Of An Unknown Salt Lab Report
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Analysis Of An Unknown Salt Lab Report
 - Setting Reading Goals Analysis Of An Unknown Salt Lab Report
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Analysis Of An Unknown Salt Lab Report
 - Fact-Checking eBook Content of Analysis Of An Unknown Salt Lab Report
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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