

Analytical Chemistry
3rd Long Exam

PART I. MULTIPLE CHOICE. Choose the letter that corresponds to the **BEST** answer. Write **[]** when the answer is **NOT** in the choices. (25 pts)

1. Transfer of electrons from one reactant to another

- a. precipitation rxn
- b. half rxn
- c. redox rxn
- d. neutralization rxn

2. What is the mole ratio of Ce(IV) to NO_2^- for the oxidation of nitrite to nitrate using Ce^{4+} in the presence of an acid:



- a. 1:2
- b. 2:1
- c. 1:1
- d. 2:2

3. How many mole(s) of electron is(are) produced in the oxidation of sodium bisulfate (NaHSO_4) to BiO^+ and H^+ under acidic conditions?

- a. 1
- b. 2
- c. 3
- d. 4

4. Azide (N_3^-) was added in the determination of the dissolved oxygen in water sample to remove interferences caused by the presence of what ion?

- a. chloride
- b. fluoride
- c. nitrite
- d. nitrate

5. Which of the following statements is(are) true?

- I. Prepared KMnO_4 solutions should be stored in amber bottles
- II. KMnO_4 solutions are boiled to remove CO_2 which makes the solution basic
- III. Mn(II) is the product of the degradation of KMnO_4
- IV. KMnO_4 solutions are filtered to remove manganese(IV) oxide

- a. I only
- b. I and II
- c. I and IV
- d. I, II and III

6. What is the oxidation state of carbon in oxalic acid?

- a. +2
- b. +4
- c. +6
- d. +7

7. The complex formed between Cu(I) and 1,10-phenanthroline has a molar absorptivity of $7000 \text{ L cm}^{-1} \text{ mol}^{-1}$ at 435 nm, the wavelength of maximum absorption. What is the absorbance of a $5.17 \times 10^{-3} \text{ M}$ solution of the complex when measured in a 1.00-cm cell at 435 nm?

- a. 0.0431
- b. 0.0432
- c. 0.432
- d. 0.431

8. In the presence of iodine, thiosulfate ion is oxidized quantitatively to tetrathionate ion. What is the chemical formula of tetrathionate ion?

- a. $\text{S}_4\text{O}_7^{2-}$
- b. $\text{S}_4\text{O}_6^{2-}$
- c. $\text{S}_4\text{O}_8^{2-}$
- d. $\text{S}_4\text{O}_5^{2-}$

9. In Winkler titration, the flocc is produced from

- a. oxidation of I^- to I_2
- b. oxidation of Mn(II) to Mn(III)
- c. reduction of I^- to I_2
- d. reduction of Mn(II) to Mn(III)

10. A component of a UV-Vis spectrophotometer that restricts the radiation being measured to a narrow band that is absorbed or emitted by the analyte.

- a. light source
- b. detector
- c. radiation filter
- d. wavelength selector

11. Component of a UV-Vis instrument that operates by blocking or absorbing all but a restricted band of radiation.

- a. light source
- b. detector
- c. radiation filter
- d. wavelength selector

12. Oxidation of VO^{2+} using permanganate or cerium as standard solution produces VO(OH)_2^{+} , what is the oxidation state of vanadium in VO(OH)_2^{+} ?

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TRAC: Trends in Analytical Chemistry U A Th Brinkman, J. G. Dorsey, J R Durig, 2013-09-24 TRAC Trends in Analytical Chemistry Volume 7 provides information pertinent to the trends in the field of analytical chemistry This book discusses a variety of topics related to analytical chemistry including biomolecular mass spectroscopy affinity chromatography electrochemical detection nucleosides and protein sequencing Organized into 63 parts encompassing 158 chapters this volume begins with an overview of the significance of quality and productivity in the analytical laboratory This text then presents a comprehensive review on alcohol dehydrogenases immobilization and applications in analysis and synthesis Other chapters consider the various tests for determining the excellence of quantitative assays available for analysts to utilize for method validation This book discusses as well the primary challenge of neuropharmacologists to relate physiological functions to the many ligand binding sites identified in brain tissue The final chapter deals with the fundamentals and applications of biosensors This book is a valuable resource for analytical chemists chemical engineers clinical chemists neuropharmacologists and scientists

Understanding Advanced Organic And Analytical Chemistry: The Learner's Approach (Revised Edition) Kim Seng Chan, Jeanne Tan, 2016-09-29 This revised edition has been updated to meet the minimum requirements of the new Singapore GCE A level syllabus that would be implemented in the year 2016 Nevertheless this book is also highly relevant to students who are studying chemistry for other examination boards In addition the authors have also included more Q A to help students better understand and appreciate the chemical concepts that they are mastering

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