#### AP Chemistry Chapter 4 Practice Exam

- The Claus reactions, shown below, are used to generate elemental sulfur from hydrogen sulfide.

$$2H_2S + 3O_2 \rightarrow 2SO_2 + 2H_2O$$
  
 $SO_2 + 2H_2S \rightarrow 3S + 2H_2O$ 

How much sulfur (in grams) is produced from 21.5 grams of O<sub>2</sub>? 21.5502 |mol 02 | 2503 | 35 | 32.0675 = 43.155

- A) 97.0 g
- B) 14.4 g
- C) 21.5 g
- none of these
- B
- Iron is produced from its ore by the reactions:

$$2C(e) + O_2(e) \rightarrow 2CO(e)$$

$$Fe_2O_3(s) + 3CO(g) \rightarrow 2Fe(s) + 3CO_3(g)$$

Closes:

How many moles of  $O_2(g)$  are needed to produce 4.6 moles of Fe(s)?

- A) 2.3 mole O<sub>2</sub> B) 3.5 mole O-
- C) 4.6 mole O<sub>2</sub>
- D) 6.9 mole O<sub>2</sub>
- E) 13.8 mole O<sub>2</sub>

- B
- Suppose the reaction Ca<sub>2</sub>(PO<sub>4</sub>)<sub>2</sub> + 3H<sub>2</sub>SO<sub>4</sub> → 3CaSO<sub>4</sub> + 2H<sub>2</sub>PO<sub>4</sub> is carried out starting with 153 g. of  $Ca_2(PO_k)_2$

and 76.8 g of  $H_2SO_4$ . How much phosphoric acid will be produced?

A) 76.7 gB) 51.1 g1  $C_{0.3}(C_{0.3}) = C_{0.3}(C_{0.3}) = C_{0.3}(C$ 

- B) 51.1 g C) 229.8 g

  - D) 115.1 c E) 96.7 g
- 7 6.89 H2504 | Ind H2504 2 H2 POY 1 3 =

1

- - SO<sub>2</sub> reacts with H<sub>2</sub>S as follows:

$$2H_2S + SO_2 \rightarrow 3S + 2H_2O$$

When 7.50 g of H<sub>2</sub>S reacts with 12.75 g of SO<sub>2</sub>, which statement applies?

- A) 6.38 g of sulfur are formed.
- B) 10.6 g of sulfur are formed.
- C) 0.0216 moles of H<sub>2</sub>S remain.
- D) 1.13 g of H<sub>2</sub>S remain.
- E) SO<sub>2</sub> is the limiting reagent.

4.6 mil Fersi 300 102 = 3.45mol02

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