# Basic Chemistry Study Guide 2

#### Hydrogen bonds

- Form between two polar molecules covalent bonds
- A hydrogen's &+ region is attracted to a &- region in another molecule.
- Form between water molecules (Where are the &+ and &- regions in a water molecule?) hydrogen bonds, at the top of the symbol
- Each hydrogen bond is very weak (can be easily formed and broken).
- There are many hydrogen bonds in water. Together, these bonds give water its unique properties that are essential for life
- One of these properties is the ability to be a solvent for other polar molecules.
- Hydrogen bonds are depicted in figures as dotted or dashed lines between two atoms.

#### 2. Solutions, solvents, and solutes

- Water is the most abundant molecule in living organisms, and 60% of human body weight is water.
- Water is located inside and outside of our cells.
- 3. Body fluids are solutions
- Define the terms solution, solvent, and solute. Solvent a substance capable of dissolving other polar molecules and ionic bonds, solutes, molecules dissolved in the solvent, solution homogeneous moture.

#### 3. How substances dissolve in water

- When a solute dissolves, it becomes surrounded by water molecules.
- 2. For a solute to dissolve, it must attract water molecules.
- What kinds of substances attract water molecules? Sugars, nucleic acids, and amino acids
- When a solute dissolves it forms hydrogen bonds with water molecules.
- Define the terms ion, cation, and anion. Ion atom with a charge, cation; positively charged ion, anions; negatively charged ion.
- Define the terms hydrophilic and hydrophobic Hydrophilic attracted to water, hydrophobic, a kaid of water
- What types of molecules are hydrophilic and what types are hydrophobic? Nonpolar hydrophobic, polar hydrophilic

# **100 Question Study Guide For Chemistry**

William H. Myers, William T. Mooney

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#### **Table of Contents 100 Question Study Guide For Chemistry**

- 1. Understanding the eBook 100 Question Study Guide For Chemistry
  - The Rise of Digital Reading 100 Question Study Guide For Chemistry
  - Advantages of eBooks Over Traditional Books
- 2. Identifying 100 Question Study Guide For Chemistry
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 100 Question Study Guide For Chemistry
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from 100 Question Study Guide For Chemistry
  - Personalized Recommendations
  - $\circ\,$  100 Question Study Guide For Chemistry User Reviews and Ratings
  - 100 Question Study Guide For Chemistry and Bestseller Lists
- 5. Accessing 100 Question Study Guide For Chemistry Free and Paid eBooks
  - 100 Question Study Guide For Chemistry Public Domain eBooks
  - 100 Question Study Guide For Chemistry eBook Subscription Services
  - 100 Question Study Guide For Chemistry Budget-Friendly Options

- 6. Navigating 100 Question Study Guide For Chemistry eBook Formats
  - o ePub, PDF, MOBI, and More
  - 100 Question Study Guide For Chemistry Compatibility with Devices
  - 100 Question Study Guide For Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of 100 Question Study Guide For Chemistry
  - Highlighting and Note-Taking 100 Question Study Guide For Chemistry
  - Interactive Elements 100 Question Study Guide For Chemistry
- 8. Staying Engaged with 100 Question Study Guide For Chemistry
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers 100 Question Study Guide For Chemistry
- 9. Balancing eBooks and Physical Books 100 Question Study Guide For Chemistry
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection 100 Question Study Guide For Chemistry
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine 100 Question Study Guide For Chemistry
  - Setting Reading Goals 100 Question Study Guide For Chemistry
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 100 Question Study Guide For Chemistry
  - Fact-Checking eBook Content of 100 Question Study Guide For Chemistry
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements

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