

Science

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(Chapter – 2) (Acids, Bases and Salts) (Exercises)
(Class 10)

Question 1:

A solution turns red litmus blue, its pH is likely to be

- (a) 1 (b) 4 (c) 5 (d) 10

Answer 1:

(d) 10

Basic solution has a pH value more than 7. Since the solution turns red litmus blue, its pH is likely to be 10.

Question 2:

A solution reacts with crushed egg-shells to give a gas that turns lime-water milky. The solution contains

- (a) NaCl (b) HCl (c) LiCl (d) KCl

Answer 2:

(b) HCl

The crushed egg-shells contains calcium carbonate [CaCO_3]. When HCl solution reacts with this crushed egg-shells and produces carbon dioxide which turns lime – water milky.



Question 3:

10 mL of a solution of NaOH is found to be completely neutralised by 8 mL of a given solution of HCl . If we take 20 mL of the same solution of NaOH , the amount HCl solution (the same solution as before) required to neutralise it will be

- (a) 4 ml (b) 8 ml (c) 12 ml (d) 16 ml

Answer 3:

(d) 16 ml

Because for the neutralisation, the ratio of the acid and base should be same.

Question 4:

Which one of the following types of medicines is used for treating indigestion?

- (a) Antibiotic (b) Analgesic (c) Antacid (d) Antiseptic

Answer 4:

(c) Antacid

Because antacid reduced the acidity (produced due to excess release of HCl) in the stomach.

Question 5:

Write word equations and then balanced equations for the reaction taking place when –

- (a) dilute sulphuric acid reacts with zinc granules.
(b) dilute hydrochloric acid reacts with magnesium ribbon.
(c) dilute sulphuric acid reacts with aluminium powder.
(d) dilute hydrochloric acid reacts with iron filings.

Answer 5:

(a) Sulphuric acid + Zinc $\longrightarrow$ Zinc sulphate + Hydrogen



(b) Hydrochloric acid + Magnesium $\longrightarrow$ Magnesium chloride + Hydrogen



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