Calculate the amount of heat needed to increase the temperature of 250 g of water from 20 °C to 46 °C.

q = m x C x ΔT q = 250 g x 4.18 J/g °C x 26°C q = 27,170 J or 27 kJ

Calculate the specific heat capacity of copper given that 204.75 J of energy raises the temperature of 15 g of copper from 25 °C to 60. °C.

q = m x C x ΔT C= q/m x ΔT C = 204.75 J / (15 g x 35 °C) C= 0.39 J/g °C

216 J of energy is required to raise the temperature of aluminum from 15 °C to 35 °C.
 Calculate the mass of aluminum. (Specific heat capacity of aluminum is 0.90 J/g °C).

q = m x C x ΔT m= q/C x ΔT m= 216 J/ (0.90 J/ g °C x 20 °C) m= 12 g

4. The initial temperature of 150 g of ethanol was 22 °C. What will be the final temperature of the ethanol if 3240 J was needed to raise the temperature of the ethanol? (Specific heat capacity of ethanol is 2.44 J/ g °C).

q = m x C x ΔT ΔT = q/m x C ΔT = 3240 J/ (150 g x 2.44 J/ g °C) ΔT = 8.85 °C Trust = 22 °C + 8.85 °C = 31 °C

5. When 435 J of heat is added to 3.4 g of olive oil at $21\,^{\circ}$ C, the temperature increases to 85 °C. What is the specific heat of the olive oil?

q = m x C x \(\Delta T\)
C= q/m x \(\Delta T\)
C = 435 \(\J/\) (3.4 g x 64 °C)
C = 2.0 \(\J/\) g °C

6. How many kilojoules of heat are absorbed when 1.00 L of water is heated from 18 °C to 85 °C? (Hint: You first need to determine the mass of the water, then calculate q in the requested unit.)

1 Specific Heat Answers

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