

Practice Test Ch 3

Stoichiometry

Name: _____ Per: _____

- Remember, this is practice - Do NOT cheat yourself of finding out what you are capable of doing. Be sure you follow the testing conditions outlined below.
- DO NOT USE A CALCULATOR. AP Chem does not allow the use of a calculator for the MC part of the exam, so it is time to start practicing without one. You may use ONLY a periodic table.
- While you should practice working as fast as possible, it is more important at this point in the course, that you practice without a calculator, even if it slows you down. Look for the "easy math" – common factors and rough estimation – do not do "long division" to try to get exact values. Remember it is a MC test, use the answers
- Mark which questions you would like to "go over" when we get to school in September.

1. Balance the following equation:



The balanced equation shows that 1.00 mole of NH_3 requires _____ mole(s) of O_2 :

- a. 0.57
 - b. 1.25
 - c. 1.33
 - d. 1.75
 - e. 3.5
2. Write a balanced equation for the combustion of acetaldehyde, CH_3CHO .
- When properly balanced, the equation indicates that _____ mole(s) of O_2 are required for each mole of CH_3CHO .
- a. 1
 - b. 2
 - c. 2.5
 - d. 3
 - e. 5
3. What is the total mass of products formed when 16 grams of CH_4 is burned with excess oxygen?
- a. 32 g
 - b. 36 g
 - c. 44 g
 - d. 62 g
 - e. 80 g
4. Write a balanced equation for the combustion of propane, C_3H_8 .

When balanced, the equation indicates that _____ moles of O_2 are required for each mole of C_3H_8 .

- a. 1.5
- b. 3
- c. 3.5
- d. 5
- e. 8

5. Balance the following equation with the SMALLEST WHOLE NUMBER COEFFICIENTS possible. Select the number that is the
- sum
- of the coefficients in the balanced equation:



- a. 3
 - b. 5
 - c. 6
 - d. 7
 - e. 8
6. Calculate the mass of hydrogen formed when 27 g of aluminum reacts with excess hydrochloric acid according to the balanced equation below.
- $$2\text{Al} + 6\text{HCl} \rightarrow 2\text{AlCl}_3 + 3\text{H}_2$$
- a. 1.5 g
 - b. 2.0 g
 - c. 3.0 g
 - d. 6.0 g
 - e. 12 g
7. How many grams of nitric acid, HNO_3 , can be prepared from the reaction of 138 g of NO_2 with 54.0 g H_2O according to the equation below?
- $$3\text{NO}_2 + \text{H}_2\text{O} \rightarrow 2\text{HNO}_3 + \text{NO}$$
- a. 92
 - b. 108
 - c. 126
 - d. 189
 - e. 279
8. Which of the following statements is true?
- I. The molar mass of CaCO_3 is 100.1 g mol^{-1} .
 - II. 50 g of CaCO_3 contains 9×10^{23} oxygen atoms.
 - III. A 200 g sample of CaCO_3 contains 2 moles of CaCO_3 .

- a. I only
- b. II only
- c. III only
- d. I and III only
- e. I, II, and III

Answer Key To Stoichiometry Test Ch 1

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