Question 1: Water at 20° C falls from a height of 854 m. If the whole energy is used in increasing the temperature, find out the final temperature. Specific heat of water is 4200 J/K kg.

Given that

$$T_1 = 20^{\circ}C = 20 + 273 = 293 \,\mathrm{K}$$

height = $h = 854 \,\mathrm{m}$

Specific heat of water = 4200 JK 'kg'

Final temperature = $T_2 = ?$

P.E = mgh — O

Solution:

we know that AQ = cm AT -

From \bigcirc and \bigcirc CHAT = WAL C AT = gh C (T2-T,) - gh T2 - T1 = gh T2 = gh + T1 $=\frac{9.8 \times 854}{4200} + 293$ $= \frac{8369.2}{4200} + 293 = 1.99 + 293$ T, = 294.99 K = 295-273 °C T, = 22°C © learningparadisehub.com

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Table of Contents Answer The Following Questions About The Thermodynamics Of

- 1. Understanding the eBook Answer The Following Questions About The Thermodynamics Of
 - The Rise of Digital Reading Answer The Following Questions About The Thermodynamics Of
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Answer The Following Questions About The Thermodynamics Of
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer The Following Questions About The Thermodynamics Of
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Answer The Following Questions About The Thermodynamics Of

- Personalized Recommendations
- Answer The Following Questions About The Thermodynamics Of User Reviews and Ratings
- Answer The Following Questions About The Thermodynamics Of and Bestseller Lists
- 5. Accessing Answer The Following Questions About The Thermodynamics Of Free and Paid eBooks
 - Answer The Following Questions About The Thermodynamics Of Public Domain eBooks
 - Answer The Following Questions About The Thermodynamics Of eBook Subscription Services
 - Answer The Following Questions About The Thermodynamics Of Budget-Friendly Options
- 6. Navigating Answer The Following Questions About The Thermodynamics Of eBook Formats
 - o ePub, PDF, MOBI, and More
 - Answer The Following Questions About The Thermodynamics Of Compatibility with Devices
 - Answer The Following Questions About The Thermodynamics Of Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer The Following Questions About The Thermodynamics Of
 - Highlighting and Note-Taking Answer The Following Questions About The Thermodynamics Of
 - Interactive Elements Answer The Following Questions About The Thermodynamics Of
- 8. Staying Engaged with Answer The Following Questions About The Thermodynamics Of
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer The Following Questions About The Thermodynamics Of
- 9. Balancing eBooks and Physical Books Answer The Following Questions About The Thermodynamics Of
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer The Following Questions About The Thermodynamics Of
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answer The Following Questions About The Thermodynamics Of
 - Setting Reading Goals Answer The Following Questions About The Thermodynamics Of
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer The Following Questions About The Thermodynamics Of

- Fact-Checking eBook Content of Answer The Following Questions About The Thermodynamics Of
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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