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Date: **October 4, 2012**

Student Exploration: Electron Configuration

Activity B: Atomic radii

Get the Gizmo ready:

- On the PERIODIC TABLE tab, select **Na** (sodium).
- Select the ELECTRON CONFIGURATION tab.



Question: How do the radii of atoms change across a period of the periodic table?

1. **Predict:** Positively charged protons in the nucleus of an atom are attracted to negatively charged electrons.

How do you think the **atomic radii** will change as electrons are added to a shell?

It will decrease

2. **Arrange:** Create a proper electron configuration for sodium. After clicking **Check**, note the **Electron configuration** and the **Atomic radius** now listed at right.

Sodium electron configuration: **1s2 2s2 2p6 3s1**

Atomic radius: **190 Pico meters**

3. **Compare:** Click **Next element**, and then add an electron to the magnesium atom. Click **check**, and record the electron configuration and atomic radius below.

Magnesium electron configuration: **1s2 2s2 2p6 3s2**

Atomic radius: **145 Pico meters**

4. **Gather data:** Create electron configurations for the next six elements. Record the electron configuration and atomic radius of each. (Note: The symbol for picometer is *pm*.)

Element	Number of electrons	Electron configuration	Atomic radius (pm)
Aluminum	13	1s2 2s2 2p6 3s2 3p1	118
Silicon	14	1s2 2s2 2p6 3s2 3p2	111
Phosphorus	15	1s2 2s2 2p6 3s2 3p3	98
Sulfur	16	1s2 2s2 2p6 3s2 3p4	88
Chlorine	17	1s2 2s2 2p6 3s2 3p5	79
Argon	18	1s2 2s2 2p6 3s2 3p6	71

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Anthony S. Fauci



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Table of Contents Answers For Student Exploration Electron Configuration Gizmo

1. Understanding the eBook Answers For Student Exploration Electron Configuration Gizmo
 - The Rise of Digital Reading Answers For Student Exploration Electron Configuration Gizmo
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers For Student Exploration Electron Configuration Gizmo
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers For Student Exploration Electron Configuration Gizmo
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers For Student Exploration Electron Configuration Gizmo
 - Personalized Recommendations
 - Answers For Student Exploration Electron Configuration Gizmo User Reviews and Ratings
 - Answers For Student Exploration Electron Configuration Gizmo and Bestseller Lists

5. Accessing Answers For Student Exploration Electron Configuration Gizmo Free and Paid eBooks
 - Answers For Student Exploration Electron Configuration Gizmo Public Domain eBooks
 - Answers For Student Exploration Electron Configuration Gizmo eBook Subscription Services
 - Answers For Student Exploration Electron Configuration Gizmo Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Answers For Student Exploration Electron Configuration Gizmo Compatibility with Devices
 - Answers For Student Exploration Electron Configuration Gizmo Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers For Student Exploration Electron Configuration Gizmo
 - Highlighting and Note-Taking Answers For Student Exploration Electron Configuration Gizmo
 - Interactive Elements Answers For Student Exploration Electron Configuration Gizmo
8. Staying Engaged with Answers For Student Exploration Electron Configuration Gizmo
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers For Student Exploration Electron Configuration Gizmo
9. Balancing eBooks and Physical Books Answers For Student Exploration Electron Configuration Gizmo
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers For Student Exploration Electron Configuration Gizmo
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answers For Student Exploration Electron Configuration Gizmo
 - Setting Reading Goals Answers For Student Exploration Electron Configuration Gizmo
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answers For Student Exploration Electron Configuration Gizmo
 - Fact-Checking eBook Content of Answers For Student Exploration Electron Configuration Gizmo
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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