

## MIXED GAS LAWS WORKSHEET - SOLUTIONS

1) How many moles of gas occupy 98 L at a pressure of 2.8 atmospheres and a temperature of 292 K?

$$n = \frac{PV}{RT} = \frac{(2.8 \text{ atm})(98 \text{ L})}{(0.0821 \text{ Latm/mol.K})(292 \text{ K})} = 11 \text{ moles of gas}$$

2) If 5.0 moles of  $O_2$  and 3.0 moles of  $N_2$  are placed in a 30.0 L tank at a temperature of  $25^\circ\text{C}$ , what will the pressure of the resulting mixture of gases be?  $25^\circ\text{C} = 298 \text{ K}$

$$O_2: P = \frac{nRT}{V} = \frac{(5.0 \text{ mol})(0.0821 \text{ Latm/mol.K})(298 \text{ K})}{(30.0 \text{ L})} = 4.1 \text{ atm}$$

$$N_2: P = \frac{nRT}{V} = \frac{(3.0 \text{ mol})(0.0821 \text{ Latm/mol.K})(298 \text{ K})}{(30.0 \text{ L})} = 2.4 \text{ atm}$$

$$P_{\text{tot}} = P_{O_2} + P_{N_2} = 4.1 \text{ atm} + 2.4 \text{ atm} = 6.5 \text{ atm}$$

Alternately you can substitute total moles of gas (8.0) as the total pressure depends on the total moles of gas – what type of gas is irrelevant

3) A balloon is filled with 35.0 L of helium in the morning when the temperature is  $20.0^\circ\text{C}$ . By noon the temperature has risen to  $45.0^\circ\text{C}$ . What is the new volume of the balloon?

$$T_1 = 20.0^\circ\text{C} = 293 \text{ K}, V_1 = 35.0 \text{ L}, T_2 = 45.0^\circ\text{C} = 318 \text{ K}, V_2 = ?$$

$$V_2 = \frac{V_1 T_2}{T_1} = \frac{(35.0 \text{ L})(318 \text{ K})}{(293 \text{ K})} = 38.0 \text{ L}$$

4) A 35 L tank of oxygen is at 315 K with an internal pressure of 190 atmospheres. How many moles of gas does the tank contain?

$$n = \frac{PV}{RT} = \frac{(190 \text{ atm})(35 \text{ L})}{(0.0821 \text{ Latm/mol.K})(315 \text{ K})} = 260 \text{ moles of gas}$$

5) A balloon that can hold 85 L of air is inflated with 3.5 moles of gas at a pressure of 1.0 atmosphere. What is the temperature in  $^\circ\text{C}$  of the balloon?

$$T = \frac{PV}{nR} = \frac{(1 \text{ atm})(85 \text{ L})}{(3.5 \text{ mol})(0.0821 \text{ Latm/mol.K})} = 296 \text{ K} = 23^\circ\text{C}$$

6)  $\text{CaCO}_3$  decomposes at  $1200^\circ\text{C}$  to form  $\text{CO}_2$  gas and  $\text{CaO}$ . If 25 L of  $\text{CO}_2$  are collected at  $1200^\circ\text{C}$ , what will the volume of this gas be after it cools to  $25^\circ\text{C}$ ?

$$T_1 = 1200^\circ\text{C} = 1473 \text{ K}, V_1 = 25 \text{ L}, T_2 = 25^\circ\text{C} = 298 \text{ K}, V_2 = ?$$

$$V_2 = \frac{V_1 T_2}{T_1} = \frac{(25 \text{ L})(298 \text{ K})}{(1473 \text{ K})} = 5.1 \text{ L}$$

7) A helium balloon with an internal pressure of 1.00 atm and a volume of 4.50 L at  $20.0^\circ\text{C}$  is released. What volume will the balloon occupy at an altitude where the pressure is 0.600 atm and the temperature is  $-20.0^\circ\text{C}$ ?

$$P_1 = 1.00 \text{ atm}, V_1 = 4.50 \text{ L}, T_1 = 20.0^\circ\text{C} = 293 \text{ K}, P_2 = 0.600 \text{ atm}, V_2 = ?, T_2 = -20.0^\circ\text{C} = 253 \text{ K}$$

$$V_2 = \frac{P_1 V_1 T_2}{T_1 P_2} = \frac{(1.00 \text{ atm})(4.50 \text{ L})(253 \text{ K})}{(293 \text{ K})(0.600 \text{ atm})} = 6.48 \text{ L}$$

# Answer Sheet To Mixed Gas Laws

**Colin Howson**



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## **Table of Contents Answer Sheet To Mixed Gas Laws**

1. Understanding the eBook Answer Sheet To Mixed Gas Laws
  - The Rise of Digital Reading Answer Sheet To Mixed Gas Laws
  - Advantages of eBooks Over Traditional Books
2. Identifying Answer Sheet To Mixed Gas Laws
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Answer Sheet To Mixed Gas Laws
  - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Sheet To Mixed Gas Laws
  - Personalized Recommendations
  - Answer Sheet To Mixed Gas Laws User Reviews and Ratings
  - Answer Sheet To Mixed Gas Laws and Bestseller Lists
5. Accessing Answer Sheet To Mixed Gas Laws Free and Paid eBooks
  - Answer Sheet To Mixed Gas Laws Public Domain eBooks
  - Answer Sheet To Mixed Gas Laws eBook Subscription Services
  - Answer Sheet To Mixed Gas Laws Budget-Friendly Options
6. Navigating Answer Sheet To Mixed Gas Laws eBook Formats

- ePub, PDF, MOBI, and More
- Answer Sheet To Mixed Gas Laws Compatibility with Devices
- Answer Sheet To Mixed Gas Laws Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Answer Sheet To Mixed Gas Laws
  - Highlighting and Note-Taking Answer Sheet To Mixed Gas Laws
  - Interactive Elements Answer Sheet To Mixed Gas Laws
- 8. Staying Engaged with Answer Sheet To Mixed Gas Laws
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Answer Sheet To Mixed Gas Laws
- 9. Balancing eBooks and Physical Books Answer Sheet To Mixed Gas Laws
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Answer Sheet To Mixed Gas Laws
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Answer Sheet To Mixed Gas Laws
  - Setting Reading Goals Answer Sheet To Mixed Gas Laws
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer Sheet To Mixed Gas Laws
  - Fact-Checking eBook Content of Answer Sheet To Mixed Gas Laws
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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