In the reaction : 2Al<sub>2</sub>O<sub>3(1)</sub> → 4Al<sub>(8)</sub> + 3O<sub>2(g)</sub>, the mole ratios are:

Al<sub>2</sub>O<sub>3</sub> to Al \_\_\_\_\_ Al<sub>2</sub>O<sub>3</sub> to O<sub>2</sub> \_\_\_\_\_ Al to O<sub>2</sub> \_\_\_\_ O<sub>2</sub> to Al \_\_\_\_\_

- 2. In the reaction: 2H₂ + O₂ → 2H₂O, the mole ratio of oxygen to water is: \_\_\_\_\_
- In the reaction: Ca + Cl<sub>2</sub> → CaCl<sub>2</sub>, the mole ratio of chlorine to calcium is: \_\_\_\_\_\_
- In the reaction: Zn + H₂SO<sub>4</sub> → ZnSO<sub>4</sub> + H₂, the mole ratio of zinc to sulfuric acid is:
- The equation for the Haber process is represented by:

 $N_{2(g)} + 3H_{2(g)} \rightarrow 2NH_{3(g)}$ 

To convert 9.0 moles of hydrogen to ammonia would require how many moles of nitrogen?

 In the equation: 2KClO<sub>3(g)</sub> → 2KCl<sub>(s)</sub> + 3O<sub>2(g)</sub>, how many moles of oxygen are produced when 3.00 moles of KClO<sub>3</sub> decompose completely?

 For the reaction: C<sub>(8)</sub> + 2H<sub>2(9)</sub> → CH<sub>4(9)</sub>, how many moles of H<sub>2</sub> are required to produce 10.0 moles of CH<sub>4</sub>?

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Scott C. Dulebohn

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