

CHAPTER 10: Reactions

Reaction Energy

SECTION 1

Write a balanced chemical equation for the following reactions in the space provided.

- For elements in their standard state, the value of ΔH_f° is _____.
- The formation and decomposition of a substance is represented by the following thermodynamic equations:



- Is the energy being added to or is it being subtracted from the decomposition?
- What is the correct sign for the enthalpy change in this decomposition reaction?

Write the correct answer in the space to the left. Show all your work in the space provided.

- _____ $\Delta H_{\text{f}}^\circ$ of a reaction is 200 kJ . Absorption of 1 mole of energy is that much less than decomposition, for?

- _____ A substance has a specific heat of $0.001 \text{ J/g}^\circ\text{C}$. If 1 mole of the substance is heated from 0°C to 100°C , it will absorb _____ kJ of energy.

- The product is a molecular form of oxygen, and the element has an enthalpy of 0 kJ/mol .

- What is the value of ΔH for this reaction?

Answers For Reaction Energy Section 1

SJ Ball



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