WORKSHEET - NET IONIC EQUATIONS

Write the molecular, total ionic, and net ionic equations for the following reactions:

1. Aqueous solutions of potassium phosphate and calcium nitrate are mixed.

ME $2 + 9PO_{4}(aq_{1}) + 3Ca(HO_{2})_{2}(aq_{2}) \rightarrow Ca_{3}(PO_{4})_{2}(a) + b KHO_{3}(aq_{2})$ THE $b K^{4} + 2FO_{4}^{3} + 3Ca^{24} + bHO_{5} \rightarrow Ca_{3}(FO_{4})_{2} + b K^{4} + bHO_{5}$ NIE $2 FO_{4}^{3} + 3Ca^{24} \rightarrow Ca_{3}(PO_{4})_{2}$.

2. Aqueous solutions of barium chloride and sodium sulfate are mixed.

ME $BaCl_{a}(aq) + Na_{5}Cu(aq) \rightarrow BaSO_{+}(a) + 2NaCl(aq)$ TIE $Ba^{2+} + 2Cl^{-} + 2Na^{+} + So_{4}^{2-} \rightarrow BaSO_{+} + 2Na^{+} + 2Cl^{-}$ NIE $Ba^{2+} + 5O_{4}^{2-} \rightarrow BaSO_{4}$

3. Magnesium is added to an aqueous solution of chromium (III) iodide.

ME $3 \text{ Mg}(6) + 2 \text{ Cr} \text{ Ta}(0q_1) \rightarrow 3 \text{ Mg} \text{ Ta}(0q_1) + 2 \text{ Cr}(6)$ TIE $3 \text{ Mg} + 2 \text{ Cr}^{3+} + 6 \text{ F} \rightarrow 3 \text{ Mg}^{2+} + 6 \text{ F} + 2 \text{ Cr}$ NIE $3 \text{ Mg} + 2 \text{ Cr}^{3+} \rightarrow 3 \text{ Mg}^{2+} + 2 \text{ Cr}$

4. Silver hydroxide is added to an aqueous phosphoric acid solution.

ME 3 Ag OH (\approx) + Ha POH ($a_{\rm H}$) \rightarrow Aga POH($a_{\rm H}$) + 3 HaO($a_{\rm H}$)

TIE 3 Ag OH + Ha POH \rightarrow Aga POH + 3 HaO

NIE same as TIE

5. Aqueous solutions of sodium fluoride and hydrochloric acid are mixed.

ME NaF(ag) + HCl(ag) \rightarrow HF(ag) + NaCl(ag)

TIE Na[†] + F⁻ + H⁺ + Cl⁻ \rightarrow HF + Na[†] + Cl⁻

NIE F⁻ + H⁺ \rightarrow HF

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