

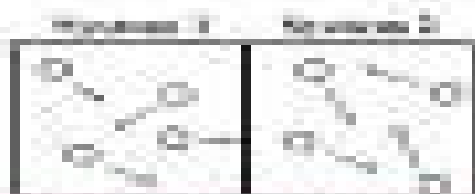
Unit 6: Thermochemistry**6.1: Heat Transfer and Thermal Equilibrium****6.2: Heat Capacity and Calorimetry****6.3: Entropy and Spontaneous Changes**

2. Use the Boltzmann-probability distribution curves below to predict temperatures of the mixtures of particles.



Temperature: is a measure of the average kinetic energy of the particles in a system. The kinetic energy for particles moving 100 m/s is 100 J , the first curve of particles with 100 J kinetic energy and 100 J of energy. The second curve is at 200 J and the third curve is at 300 J . The curves are labeled T_1 , T_2 , and T_3 from left to right. The curves shift to the right and broaden as temperature increases from T_1 to T_3 .

3. Sketch a curve to represent the relative distribution of the particles in the following two gases, samples of the same gas.



- a. In System 1, it is a higher or lower temperature than System 2? Justify your answer.
System 1 is at a higher temperature than System 2. Temperature is a measure of the average kinetic energy of the particles in a system. In System 1, the particles are at a higher energy state, and the distribution curve is shifted to the right. In System 2, the particles are at a lower energy state, and the distribution curve is shifted to the left. The curves are labeled T_1 and T_2 from left to right. The curves shift to the right and broaden as temperature increases from T_1 to T_2 .
- b. In which direction will thermal energy be transferred? Justify your answer in terms of molecular collisions.
Thermal energy will be transferred from System 1 to System 2. The particles in System 1 are at a higher energy state and have a higher average kinetic energy. The particles in System 2 are at a lower energy state and have a lower average kinetic energy. When the two systems are in contact, the particles will collide, and the energy will be transferred from the higher energy state to the lower energy state.

Answer Key Unit 6 Thermochemistry

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