

# SCH4U Unit 4 - Quiz – Equilibrium

Player: \_\_\_\_\_

Period: \_\_\_\_\_

Q1: (10)

Q2: (10)

Q3: (10)

Q4: (10)

Q5: (10)

Write the formulas for:

$$K_{eq} = \frac{[products]}{[reactants]} \quad ; \quad K_{eq} = \frac{P_{products}}{P_{reactants}}$$

$$K_{eq} = \frac{P_{CO_2}^2 P_{H_2O}}{P_{CH_4}^2 P_{O_2}} \quad ; \quad K_{eq} = \frac{P_{CO_2}^2 P_{H_2O}}{P_{CH_4}^2 P_{O_2}}$$

gpt (aq) + h<sub>2</sub>o (l) ⇌ gpt(a) + h<sub>2</sub>g(g) ⇌ gpt(a) + h<sub>2</sub>g(g)

h<sub>2</sub>o(l) ⇌ h<sub>2</sub>o(g)

## Question 1:

1. Hydrogen sulfide, H<sub>2</sub>S, is a slightly poisonous gas. At 100°C, an equilibrium mixture contains 0.020 moles of hydrogen, H<sub>2</sub>, 0.010 moles of hydrogen sulfide, and an undetermined amount of sulfur, S<sub>2</sub>. The reaction is:



If the volume of the mixture is 1.0 L, what amount of S<sub>2</sub> is present at equilibrium?

a. 0.010

b. 0.020 moles of S<sub>2</sub>

c. 0.010 moles of S<sub>2</sub>

d. 0.020 moles of S<sub>2</sub>

2. Determine the value of K<sub>eq</sub> at equilibrium.

3. Write the:

4. Write the chemical formula for the products of the reaction above.

5. Calculate:

|                 | 2H <sub>2</sub> (g) | S <sub>2</sub> (g) | 2H <sub>2</sub> S(g) |
|-----------------|---------------------|--------------------|----------------------|
| Initial (moles) |                     |                    |                      |
| Change (moles)  |                     |                    |                      |
| Final (moles)   | 0.020               | 0.010              |                      |

6. The balanced equation:

# Answers For Sch4u Unit 4

**H Kauffman**



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