

LAB MANUAL CHEMISTRY CLASS XII

Easy!

Experiment 1

Aim: To prepare 250 ml of 0.1N of Mohr's salt

Theory

Molecular mass of Mohr's salt

$\text{FeSO}_4 \cdot (\text{NH}_4)_2\text{SO}_4 \cdot 6\text{H}_2\text{O}$ is 392 g

Hence, for preparing 250 ml of 0.1N Mohr's salt

solution 3.92 g of salt is required

for preparing 250 ml of 0.1N Mohr's salt
solution $\frac{3.92 \times 250 \times 100}{1000} = 9.9 \text{ g}$ is required

Apparatus required: watch glass, weights, 250 ml beaker, glass rod, wash bottle.

Chemicals required: Mohr's salt, conc. H_2SO_4 ,
distilled water.

Procedure

1. Record weight of clean dry watch glass

2. Weigh and record exactly 9.9 g of
Mohr's salt crystals

Answer To Hands On Lab Manual Chemistry

Huangqi Zhang



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Laboratory Manual for Principles of General Chemistry Jo Allan Beran, 2010-11-01 This new edition of the Beran lab manual emphasizes chemical principles as well as techniques The manual helps students understand the timing and situations for the various techniques The Beran lab manual has long been a market leading lab manual for general chemistry Each experiment is presented with concise objectives a comprehensive list of techniques and detailed lab intros and step by step procedures

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