

Bachelor of Engineering Examination, 2023

(1st Year, 1st Semester)

CHEMISTRY

Time: 3 hours

Full Marks: 100

Use separate answer sheet for Chemistry-I, Chemistry-II and Chemistry-III

CO-1

CHEMISTRY-I

Answer any three questions from CO-1

Q1. (a) Derive the expression for the calculation of the radius of Bohr's orbit and the velocity of the electron in hydrogen atom. 2+2

(b) What would be the wavelength of the 2nd line in the Lyman series for the H atom [Given $R_H = 109678 \text{ cm}^{-1}$]. 2

(c) Calculate the de Broglie wavelength of a bullet having mass = $2 \times 10^{-3} \text{ kg}$, moving with the speed of 300 ms^{-1} . 2

(d) State and explain "Heisenberg uncertainty principle". 2

Q2. (a) What are the merits and demerits of Bohr's atomic model? 3

(b) Calculate the ionization energy of Li^{2+} ion. 1.5

(c) Define lattice energy. Construct Born-Haber cycle of NaCl. 1 + 2.5

(d) Applying Fajan's rule explain the trend on thermal stabilities of the following compounds

. BeCO_3 , MgCO_3 , CaCO_3 , SrCO_3 , BaCO_3 . 2

Q3. (a) Define buffer solutions with examples. 2

(b) Discuss the mechanism of buffer action. 3

(c) Determine the pH of a solution obtained by mixing equal volumes of 0.1 M NH_4Cl and 0.02 M NH_4OH . [Given that K_b for NH_4OH is 1.8×10^{-5}] 5

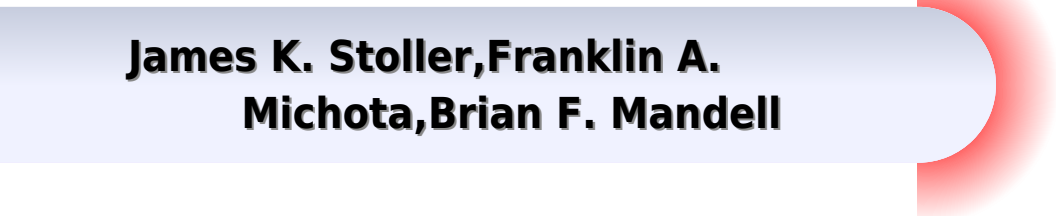
Q3. (a) Explain the trend of H-X-H angles (X = central atom) in CH_4 , NH_3 and H_2O .

3

[Turn over

Answer To 1st Semester Chemistry Exam Review

**James K. Stoller, Franklin A.
Michota, Brian F. Mandell**



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Table of Contents Answer To 1st Semester Chemistry Exam Review

1. Understanding the eBook Answer To 1st Semester Chemistry Exam Review
 - The Rise of Digital Reading Answer To 1st Semester Chemistry Exam Review
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer To 1st Semester Chemistry Exam Review
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer To 1st Semester Chemistry Exam Review
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer To 1st Semester Chemistry Exam Review

- Personalized Recommendations
- Answer To 1st Semester Chemistry Exam Review User Reviews and Ratings
- Answer To 1st Semester Chemistry Exam Review and Bestseller Lists
- 5. Accessing Answer To 1st Semester Chemistry Exam Review Free and Paid eBooks
 - Answer To 1st Semester Chemistry Exam Review Public Domain eBooks
 - Answer To 1st Semester Chemistry Exam Review eBook Subscription Services
 - Answer To 1st Semester Chemistry Exam Review Budget-Friendly Options
- 6. Navigating Answer To 1st Semester Chemistry Exam Review eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer To 1st Semester Chemistry Exam Review Compatibility with Devices
 - Answer To 1st Semester Chemistry Exam Review Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer To 1st Semester Chemistry Exam Review
 - Highlighting and Note-Taking Answer To 1st Semester Chemistry Exam Review
 - Interactive Elements Answer To 1st Semester Chemistry Exam Review
- 8. Staying Engaged with Answer To 1st Semester Chemistry Exam Review
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer To 1st Semester Chemistry Exam Review
- 9. Balancing eBooks and Physical Books Answer To 1st Semester Chemistry Exam Review
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer To 1st Semester Chemistry Exam Review
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answer To 1st Semester Chemistry Exam Review
 - Setting Reading Goals Answer To 1st Semester Chemistry Exam Review
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer To 1st Semester Chemistry Exam Review

- Fact-Checking eBook Content of Answer To 1st Semester Chemistry Exam Review
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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