

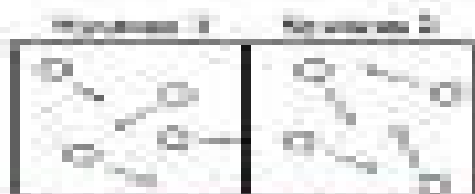
Unit 6: Thermochemistry**6.1: Heat Transfer and Thermal Equilibrium****6.2: Heat Capacity and Calorimetry****6.3: Entropy and Spontaneous Changes**

2. Use the Boltzmann probability distribution curves below to predict temperatures of the mixtures of particles.



Temperature: is a measure of the average kinetic energy of the particles in a system. The kinetic energy for particles moving 100 m/s is 100 J , the first curve of particles with 100 J of energy will change, resulting in a lower rate of temperature increase. The Boltzmann distribution curve shows that the average velocity of each of the distributions is increasing (less to large) and the average temperature is increasing.

3. Sketch a curve used to represent the relative distribution of the particles in the following two gas samples of the same gas.



- a. In System 1, it is a higher or lower temperature than System 2? Justify your answer.
 System 1 is at a lower temperature than System 2. Temperature is a measure of the average KE of the particles in the system. In System 1, the particles are more closely packed, and all the intermolecular interactions have a higher energy. According to the equation $KE = \frac{1}{2}mv^2$, the particles all of the particles have the same mass, velocity is the same magnitude of all particles KE. In the particles in System 1 are smaller, which particles are moving at slower velocities and therefore a lower average KE.
- b. Are there different total kinetic energy for System 2? Justify your answer in terms of molecular collisions.
 All particles in System 1 and System 2 are at the same temperature. Temperature represents the average kinetic energy of the particles in the system. The particles in System 2 are more spread out, and the intermolecular interactions are less. The particles in System 2 are moving at higher velocities and therefore a higher average KE.

Answer Key Unit 6 Thermochemistry

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