

OpenQuest 13

Orbital Configuration

Name _____

Date _____

Class _____

Information: Energy of Sublevels

Not all orbitals have the same energy of energy. For example, orbitals in the $3p$ sublevel have more energy than orbitals in the $3s$ sublevel. The following list shows orbitals ordered from lowest to highest energy.

$1s, 2s, 2p, 3s, 3p, 4s, 3d, 4p, 5s, 4d, 5p, 6s, 4f, 5d, 6p, 7s, 5f, 6d, 7p, 8s, 7d, 8p, 9s, 7f, 8d, 9p, \dots$

For help you have information for which orbitals included. Here's the orbital configuration for Boron.

$1s^2, 2s^2, 2p^1, 3s^0, 3p^0, 4s^0, 3d^0, 4p^0, 5s^0, 4d^0, 5p^0, 6s^0, 4f^0, 5d^0, 6p^0, 7s^0, 5f^0, 6d^0, 7p^0, 8s^0, 7d^0, 8p^0, 9s^0, 7f^0, 8d^0, 9p^0, \dots$

Not the 4 orbitals appear in the list of orbitals. This is because they have more energy. For example, the 4s orbital has higher energy than the 3d orbital, so the 4s orbital has a higher energy than the 3d orbital.

When electrons occupy orbitals, they try to have the lowest possible energy possible. (This is called the **lowest energy principle**.) An electron will occupy a 3d orbital before the 4s orbital is filled, and a 4s orbital will be filled before the 3d orbital is filled. Additionally, an electron will fill a sublevel. (This is called the **Pauli Exclusion Principle**.) When two electrons occupy the same orbital, they must spin in opposite directions—orbital spins are opposite directions.

Critical Thinking Questions

- a. How many electrons would it take to fill a sublevel before beginning to fill the 4s orbital? **10 electrons (the 4s orbital is filled, all the orbitals are filled, and the 4s orbital is filled).**
b. What is the first element that has enough electrons to fill the 4s orbital? **Calcium (Ca).**
c. What is the first element that has enough electrons to fill the 4s orbital? **Calcium (Ca).**
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b. What is the first element that has enough electrons to fill the 4s orbital? **Calcium (Ca).**
c. What is the first element that has enough electrons to fill the 4s orbital? **Calcium (Ca).**

Information: Hund's Rule

Electrons prefer to "split" or "spread". Electrons in the same orbital will have opposite spins. Electrons in the same orbital will have opposite spins. Electrons in the same orbital will have opposite spins. Electrons in the same orbital will have opposite spins.



If an orbital is empty, an electron will go into the orbital. The first electron will go into the orbital. The first electron will go into the orbital. The first electron will go into the orbital.



Hund's rule says that electrons in the same orbital will have opposite spins. Electrons in the same orbital will have opposite spins. Electrons in the same orbital will have opposite spins. Electrons in the same orbital will have opposite spins.

The following are the electron orbital diagrams for the elements, nitrogen and oxygen. Under the diagram, the number of electrons is given to show Hund's rule. The 2p orbitals are filled, beginning with nitrogen's configuration.



Critical Thinking Questions

- How many "paired" electrons are in a nitrogen atom? **2**
- Why does oxygen's electron distribution is given as $1s^2, 2s^2, 2p^4$? Why not $1s^2, 2s^2, 2p^4$? **The 2p orbitals are already full, and the 2p orbitals are already full.**
- Why do electrons in nitrogen have opposite spins? **1s, 2s, 2p, 3s, 3p, 4s, 4p, 5s, 5p, 6s, 6p, 7s, 7p, 8s, 8p, 9s, 9p, 10s, 10p, 11s, 11p, 12s, 12p, 13s, 13p, 14s, 14p, 15s, 15p, 16s, 16p, 17s, 17p, 18s, 18p, 19s, 19p, 20s, 20p, 21s, 21p, 22s, 22p, 23s, 23p, 24s, 24p, 25s, 25p, 26s, 26p, 27s, 27p, 28s, 28p, 29s, 29p, 30s, 30p, 31s, 31p, 32s, 32p, 33s, 33p, 34s, 34p, 35s, 35p, 36s, 36p, 37s, 37p, 38s, 38p, 39s, 39p, 40s, 40p, 41s, 41p, 42s, 42p, 43s, 43p, 44s, 44p, 45s, 45p, 46s, 46p, 47s, 47p, 48s, 48p, 49s, 49p, 50s, 50p, 51s, 51p, 52s, 52p, 53s, 53p, 54s, 54p, 55s, 55p, 56s, 56p, 57s, 57p, 58s, 58p, 59s, 59p, 60s, 60p, 61s, 61p, 62s, 62p, 63s, 63p, 64s, 64p, 65s, 65p, 66s, 66p, 67s, 67p, 68s, 68p, 69s, 69p, 70s, 70p, 71s, 71p, 72s, 72p, 73s, 73p, 74s, 74p, 75s, 75p, 76s, 76p, 77s, 77p, 78s, 78p, 79s, 79p, 80s, 80p, 81s, 81p, 82s, 82p, 83s, 83p, 84s, 84p, 85s, 85p, 86s, 86p, 87s, 87p, 88s, 88p, 89s, 89p, 90s, 90p, 91s, 91p, 92s, 92p, 93s, 93p, 94s, 94p, 95s, 95p, 96s, 96p, 97s, 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Answers For Chemquest

**Corwin Hansch, Peter George
Sammes, John Bodenhan Taylor**

Answers For Chemquest :

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