

# Chemistry - Review Chapters (415) 6

Directions: Answer each of the following questions. You need not use complete sentences.

1. By what property did Mendeleev arrange the elements?

**ATOMIC MASS**

2. By what property did Moseley suggest that the periodic table be arranged?

**ATOMIC NUMBER**

3. What is a period? How many are there in the periodic table?

**A HORIZONTAL ROW IN THE PERIODIC TABLE; 7**

4. What is a group (also called a family)? How many are there in the periodic table?

**A VERTICAL COLUMN IN THE PERIODIC TABLE; 18**

5. State the number of valence electrons in an atom of:

|           |            |             |            |
|-----------|------------|-------------|------------|
| a. sulfur | b. calcium | c. chlorine | d. arsenic |
| <b>6</b>  | <b>2</b>   | <b>7</b>    | <b>5</b>   |

6. Give the names and chemical symbols for the elements that correspond to these atomic numbers:

|                 |                  |                    |                    |
|-----------------|------------------|--------------------|--------------------|
| a. 10           | b. 18            | c. 36              | d. 90              |
| <b>Ne, NEON</b> | <b>Ar, ARGON</b> | <b>Kr, KRYPTON</b> | <b>Th, THORIUM</b> |

7. List, by number, both the period and group of each of these elements.

|              | <u>Symbol</u> | <u>Period</u> | <u>Group</u> |
|--------------|---------------|---------------|--------------|
| a. beryllium | Be            | <b>2</b>      | <b>2</b>     |
| b. iron      | Fe            | <b>4</b>      | <b>8</b>     |
| c. lead      | Pb            | <b>6</b>      | <b>14</b>    |

8. Which of the following pairs of elements belong to the same period?

|                     |              |              |              |
|---------------------|--------------|--------------|--------------|
| <b>a. Na and Cl</b> | b. Na and Li | c. Na and Cu | d. Na and Ne |
|---------------------|--------------|--------------|--------------|

9. Which of the following pairs of elements belong to the same group?

|             |              |                    |              |
|-------------|--------------|--------------------|--------------|
| a. H and He | b. Li and Be | <b>c. C and Pb</b> | d. Ga and Ge |
|-------------|--------------|--------------------|--------------|

10. How does an element's period number relate to the number of the energy level of its valence electrons?

**PERIOD NUMBER = ENERGY LEVEL OF VALENCE ELECTRONS**

11. What are the transition elements?

**GROUPS 3-12**

12. In what type of orbitals are the actinide and lanthanide electrons found?

**f ORBITALS**

# Answers Prentice Hall Chemistry Ch6 Chemical Names And Formulas

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