



Name:

Date:

Student Exploration: Phase Changes (4b)

Directions: Follow the instructions to go through the simulation. Respond to the questions and prompts in the orange boxes.

Vocabulary: altitude, boil, boiling point, freeze, freezing point, gas, liquid, melt, melting point, phase, solid

Activity B:

Temperature and molecular motion

Get the Gizmo ready:

- Click **Reset**, and select the **Micro view**.
- Set **Ice volume** to 0 cm³.
- Set **Add/remove heat energy** to 0 J/s.



Question: Why do phase changes occur?

1. **Compare:** Set the **Water temperature** to 0 °C and click **Play**. Observe the water molecules. Click **Reset**, set the **Water temperature** to 100 °C, and click **Play** again.

What do you notice?

There is no difference.

2. **Observe:** Click **Reset**. The **mean molecular speed** of the water molecules is displayed below the container. Set the **Water temperature** to 0 °C and **Add/remove heat energy** to 400 J/s. Click **Play**.

A. How does the mean speed of the water molecules change as they are heated?

The water gets hotter (has more heat energy).

B. Does the mean molecular speed change as much as the temperature as the water heats up? Explain.

Yes, the more the temperature increases the more the molecules of the water speed up (until they hit the boiling point, where they cannot get hotter any longer.)

3. **Explain:** How is temperature related to the motions of molecules?

The higher the temperature, the faster the molecules move.

4. **Observe:** Click **Reset**. Set the **Water temperature** to 20 °C and the **Ice volume** to 50 cm³. Set **Add/remove heat energy** to 0 J/s. Click **Play**. How do the molecules in the liquid interact with the molecules in the solid?

Answers For Activity C Gizmo Phase Changes

M Carnoy

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name M Carnoy.

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Table of Contents Answers For Activity C Gizmo Phase Changes

1. Understanding the eBook Answers For Activity C Gizmo Phase Changes
 - The Rise of Digital Reading Answers For Activity C Gizmo Phase Changes
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers For Activity C Gizmo Phase Changes
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers For Activity C Gizmo Phase Changes
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers For Activity C Gizmo Phase Changes
 - Personalized Recommendations
 - Answers For Activity C Gizmo Phase Changes User Reviews and Ratings
 - Answers For Activity C Gizmo Phase Changes and Bestseller Lists
5. Accessing Answers For Activity C Gizmo Phase Changes Free and Paid eBooks

- Answers For Activity C Gizmo Phase Changes Public Domain eBooks
- Answers For Activity C Gizmo Phase Changes eBook Subscription Services
- Answers For Activity C Gizmo Phase Changes Budget-Friendly Options
- 6. Navigating Answers For Activity C Gizmo Phase Changes eBook Formats
 - ePub, PDF, MOBI, and More
 - Answers For Activity C Gizmo Phase Changes Compatibility with Devices
 - Answers For Activity C Gizmo Phase Changes Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers For Activity C Gizmo Phase Changes
 - Highlighting and Note-Taking Answers For Activity C Gizmo Phase Changes
 - Interactive Elements Answers For Activity C Gizmo Phase Changes
- 8. Staying Engaged with Answers For Activity C Gizmo Phase Changes
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers For Activity C Gizmo Phase Changes
- 9. Balancing eBooks and Physical Books Answers For Activity C Gizmo Phase Changes
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers For Activity C Gizmo Phase Changes
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answers For Activity C Gizmo Phase Changes
 - Setting Reading Goals Answers For Activity C Gizmo Phase Changes
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answers For Activity C Gizmo Phase Changes
 - Fact-Checking eBook Content of Answers For Activity C Gizmo Phase Changes
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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