

AP Chemistry Chapter 4 Practice Exam

D

1. The Claus reactions, shown below, are used to generate elemental sulfur from hydrogen sulfide.

How much sulfur (in grams) is produced from 21.5 grams of O_2 ?

A) 97.0 g

B) 14.4 g

C) 21.5 g

D) 43.1 g

E) none of these

$$\frac{21.5 \text{ g O}_2}{32.00 \text{ g O}_2} \times \frac{1 \text{ mol O}_2}{1 \text{ mol O}_2} \times \frac{2 \text{ SO}_2}{3 \text{ O}_2} \times \frac{3 \text{ S}}{\text{SO}_2} \times \frac{32.06 \text{ g S}}{1 \text{ mol S}} = 43.1 \text{ g S}$$

B

2. Iron is produced from its ore by the reactions:

How many moles of $\text{O}_2(g)$ are needed to produce 4.6 moles of $\text{Fe}(s)$?A) 2.3 mole O_2 B) 3.5 mole O_2 C) 4.6 mole O_2 D) 6.9 mole O_2 E) 13.8 mole O_2

$$\frac{4.6 \text{ mol Fe}(s)}{2 \text{ Fe}} \times \frac{3 \text{ CO}}{2 \text{ Fe}} \times \frac{1 \text{ O}_2}{2 \text{ CO}} = 3.45 \text{ mol O}_2$$

B

3. Suppose the reaction
- $\text{Ca}_3(\text{PO}_4)_2 + 3\text{H}_2\text{SO}_4 \rightarrow 3\text{CaSO}_4 + 2\text{H}_3\text{PO}_4$
- is carried out starting with 153 g of
- $\text{Ca}_3(\text{PO}_4)_2$

and 76.8 g of H_2SO_4 . How much phosphoric acid will be produced?

A) 76.7 g

B) 31.1 g

C) 229.8 g

D) 115.1 g

E) 96.7 g

$$\frac{153 \text{ g Ca}_3(\text{PO}_4)_2}{310 \text{ g Ca}_3(\text{PO}_4)_2} \times \frac{1 \text{ mol Ca}_3(\text{PO}_4)_2}{1 \text{ mol Ca}_3(\text{PO}_4)_2} \times \frac{2 \text{ H}_3\text{PO}_4}{1 \text{ Ca}_3(\text{PO}_4)_2} = 0.97 \text{ mol H}_3\text{PO}_4$$

$$\frac{76.8 \text{ g H}_2\text{SO}_4}{98.08 \text{ g H}_2\text{SO}_4} \times \frac{1 \text{ mol H}_2\text{SO}_4}{1 \text{ mol H}_2\text{SO}_4} \times \frac{2 \text{ H}_3\text{PO}_4}{3 \text{ H}_2\text{SO}_4} = 0.51 \text{ mol H}_3\text{PO}_4$$

B

- 4.
- SO_2
- reacts with
- H_2S
- as follows:

When 7.50 g of H_2S reacts with 12.75 g of SO_2 , which statement applies?

A) 6.38 g of sulfur are formed.

B) 10.6 g of sulfur are formed.C) 0.0216 moles of H_2S remain.D) 1.13 g of H_2S remain.E) SO_2 is the limiting reagent.

$$\frac{7.50 \text{ g H}_2\text{S}}{34.03 \text{ g H}_2\text{S}} \times \frac{1 \text{ mol H}_2\text{S}}{1 \text{ mol H}_2\text{S}} \times \frac{3 \text{ S}}{2 \text{ H}_2\text{S}} \times \frac{32.06 \text{ g S}}{1 \text{ mol S}} = 10.89 \text{ g S}$$

$$\frac{12.75 \text{ g SO}_2}{64.06 \text{ g SO}_2} \times \frac{1 \text{ mol SO}_2}{1 \text{ mol SO}_2} \times \frac{3 \text{ S}}{1 \text{ SO}_2} \times \frac{32.06 \text{ g S}}{1 \text{ mol S}} = 19.13 \text{ g S}$$

Answers For Ap Chemistry Test Open Ended

**Randy E. Bennett, Linda Darling-
Hammond, Aneesha Badrinarayan**



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Children's Reading Comprehension and Assessment Scott G. Paris, Steven A. Stahl, 2005-03-23 The volume identifies critical issues and cutting edge research in assessing children's reading comprehension

Socioculturally Responsive Assessment Randy E. Bennett, Linda Darling-Hammond, Aneesh Badrinarayan, 2025-05-01 Socioculturally Responsive Assessment assembles the best available thinking from within and outside the educational measurement community about the theoretical foundations and systems level policy implications of formal assessment programs designed to be socioculturally responsive Synthesized from culturally responsive assessment design and practices culturally relevant pedagogy and funds of knowledge universal design for learning the learning sciences and other literatures this emerging concept affirms that students learning and performance is inextricably tied to the social cultural and linguistic contexts in which they live and develop knowledge Across four sections this book provides an argument and initial evidence for impact on students users and assessment quality offers guidance for implementation and examines the potential limitations pitfalls barriers and measurement issues that such programs will inevitably raise Scholars teaching faculty test developers and policymakers will come away with integral foundations new assessment approaches and a greater sense of the potential for positive impact that these assessments may afford

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