

Student Exploration: Calorimetry Lab

Gizmo Warm-up

A calorimeter is an insulated container filled with a liquid, usually water. When a hot object is placed in the calorimeter, heat energy is transferred from the object to the water and the water heats up.

Calorimeters can be used to find a substance's **specific heat capacity**. You will use the Calorimetry Lab Gizmo™ to determine the specific heat capacities of various substances.



- On the **SIMULATION** pane, select **Copper**. Use the slider to set its **Mass** to 200 g. Set the **Water mass** to 200 g. Check that the **Water temp** is set to 30.0 °C and the copper's **Temp** is 90 °C. Select the **GRAPH** tab, and click **Play** (▶).

A. What was the **Final temperature** of the copper and the water?

34.96°C

B. How much did the temperature of the copper change?

$$\begin{aligned} T &= T_f - T_i \\ &= (34.96^\circ\text{C} - 90^\circ\text{C}) \\ &= -55.04^\circ\text{C} \end{aligned}$$

The temperature of the copper changed about 55°C.

C. How much did the temperature of the water change?

$$\begin{aligned} T &= T_f - T_i \\ &= (34.96^\circ\text{C} - 30^\circ\text{C}) \\ &= 4.96^\circ\text{C} \end{aligned}$$

The temperature of the water changed about 5°C.

- Specific heat capacity can be described as a substance's resistance to temperature changes. Which substance has a greater specific heat capacity, copper or water? Explain.

$$\begin{aligned} c_w &= 4.184 \text{ J/g} \\ c_c &= 0.385 \text{ J/g} \end{aligned}$$

Water has a greater specific heat capacity.

Activity A:

Heat transfer

Get the Gizmo ready:

- Click **Reset** (↺).



Question: What factors determine how heat energy transfers between objects?

- Predict: In the Gizmo warm-up, you saw how 200 g of 90 °C copper transfers heat to 200 g of 30.0 °C water.

A. How do you think increasing the water's mass would affect the final temperature?

Answers For Calorimetry Gizmo

LL Leslie

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'LL Leslie' text.

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Table of Contents Answers For Calorimetry Gizmo

1. Understanding the eBook Answers For Calorimetry Gizmo
 - The Rise of Digital Reading Answers For Calorimetry Gizmo
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers For Calorimetry Gizmo
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers For Calorimetry Gizmo
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers For Calorimetry Gizmo
 - Personalized Recommendations
 - Answers For Calorimetry Gizmo User Reviews and Ratings

- Answers For Calorimetry Gizmo and Bestseller Lists
- 5. Accessing Answers For Calorimetry Gizmo Free and Paid eBooks
 - Answers For Calorimetry Gizmo Public Domain eBooks
 - Answers For Calorimetry Gizmo eBook Subscription Services
 - Answers For Calorimetry Gizmo Budget-Friendly Options
- 6. Navigating Answers For Calorimetry Gizmo eBook Formats
 - ePub, PDF, MOBI, and More
 - Answers For Calorimetry Gizmo Compatibility with Devices
 - Answers For Calorimetry Gizmo Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers For Calorimetry Gizmo
 - Highlighting and Note-Taking Answers For Calorimetry Gizmo
 - Interactive Elements Answers For Calorimetry Gizmo
- 8. Staying Engaged with Answers For Calorimetry Gizmo
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers For Calorimetry Gizmo
- 9. Balancing eBooks and Physical Books Answers For Calorimetry Gizmo
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers For Calorimetry Gizmo
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answers For Calorimetry Gizmo
 - Setting Reading Goals Answers For Calorimetry Gizmo
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answers For Calorimetry Gizmo
 - Fact-Checking eBook Content of Answers For Calorimetry Gizmo
 - Distinguishing Credible Sources

-
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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