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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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$$8 3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} = 8 3 \text{ mole H}_2\text{O}$$

 8 3 mole H₂O $\times$ 18 g H₂O / mole H₂O = 149 4 g H₂O is your theoretical yield so percent yield = $\frac{137 1 \text{ g H}_2\text{O}}{149 4 \text{ g H}_2\text{O}} \times 100 = 91 77\%$
 problems 6 8 are similar to problem 5

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$$\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$$

 what is the limiting reagent if 8 52 g are actually formed what is the

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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$

 if the reaction of 21 3 grams of O₂ produces a 47 3 % yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have

$$\frac{21 3 \text{ g O}_2}{32 \text{ g O}_2/\text{mole}} = 0 666 \text{ moles O}_2$$

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web a write a balanced equation for the reaction

$$\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$$

 b determine the theoretical yield of Al₂O₃ 160 0

g al x 1 mol al x 2 mol al₂O₃ x 101 964 g al₂O₃ 302 3 g al₂O₃ 26 982 g al 4 mol al 1 mol al₂O₃ c determine the percent yield
 yield actual x 100 260 0 g x 100 86 01 theoretical 302 3 g

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 weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12
 c is the limiting reagent 4 33 g of h 2 are left over 13

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web the percent yield is calculated as follows text percent yield $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ percent
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 what is the

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80 g I_2O_5 1 mol I_2O_5 1 mol I_2 $\times$ 1 333.8 g I_2 1 mol I_2O_5 28 g CO 1 mol CO

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