

Practicing the Mole - - Even Problems Answer Key

Calculate the mass in grams of each of the following:

2. 8.00 moles of aluminum

$$\frac{8.00 \text{ mol} \mid 26.98 \text{ g}}{1 \text{ mol}} = 216 \text{ g Al}$$

6. 7.00 moles of iodine (I₂)

$$\frac{7.00 \text{ mol} \mid 253.8 \text{ g}}{1 \text{ mol}} = 1,780 \text{ g I}_2$$

4. 2.00×10^2 moles of chlorine (Cl₂)

$$\frac{200. \text{ mol} \mid 70.90 \text{ g}}{1 \text{ mol}} = 14,200 \text{ g Cl}_2$$

8. 9.20 moles of iron

$$\frac{9.20 \text{ mol} \mid 55.85 \text{ g}}{1 \text{ mol}} = 514 \text{ g Fe}$$

Calculate the number of moles of atoms in each of the following:

10. 281 g of iron

$$\frac{281 \text{ g} \mid 1 \text{ mol}}{55.85 \text{ g}} = 5.03 \text{ mol Fe}$$

14. 93.0 g of phosphorus

$$\frac{93.0 \text{ g} \mid 1 \text{ mol}}{30.97 \text{ g}} = 3.00 \text{ mol P}$$

12. 72.0 g of magnesium

$$\frac{72.0 \text{ g} \mid 1 \text{ mol}}{24.31 \text{ g}} = 2.96 \text{ mol Mg}$$

16. 1035 g of lead

$$\frac{1035 \text{ g} \mid 1 \text{ mol}}{207.2 \text{ g}} = 4.995 \text{ mol Pb}$$

Calculate the number of atoms in each of the following:

18. 5.40 moles of iron

$$\frac{5.40 \text{ mol} \mid 6.022 \times 10^{23} \text{ atoms}}{1 \text{ mole}} = 3.25 \times 10^{24} \text{ atoms Fe}$$

20. 10.0 moles of titanium

$$\frac{10.0 \text{ mol} \mid 6.022 \times 10^{23} \text{ atoms}}{1 \text{ mole}} = 6.02 \times 10^{24} \text{ atoms Ti}$$

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Karen C. Timberlake



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Table of Contents Answer Questions And Problems Timberlake Chemistry

1. Understanding the eBook Answer Questions And Problems Timberlake Chemistry
 - The Rise of Digital Reading Answer Questions And Problems Timberlake Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Questions And Problems Timberlake Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Questions And Problems Timberlake Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Questions And Problems Timberlake Chemistry

- Personalized Recommendations
- Answer Questions And Problems Timberlake Chemistry User Reviews and Ratings
- Answer Questions And Problems Timberlake Chemistry and Bestseller Lists
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 - Answer Questions And Problems Timberlake Chemistry Enhanced eBook Features
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 - Interactive Elements Answer Questions And Problems Timberlake Chemistry
- 8. Staying Engaged with Answer Questions And Problems Timberlake Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Questions And Problems Timberlake Chemistry
- 9. Balancing eBooks and Physical Books Answer Questions And Problems Timberlake Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Questions And Problems Timberlake Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answer Questions And Problems Timberlake Chemistry
 - Setting Reading Goals Answer Questions And Problems Timberlake Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer Questions And Problems Timberlake Chemistry

- Fact-Checking eBook Content of Answer Questions And Problems Timberlake Chemistry
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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