

Gravimetric Analysis of Chloride in Solution

Experiment 10
Lab 10
Date: _____

Introduction:

This experiment involves the gravimetric analysis of a solution containing chloride ions. The principle of gravimetric analysis is based on the fact that the mass of a substance is proportional to the number of moles of the substance. In this experiment, the chloride ions in the solution are precipitated as silver chloride (AgCl) by adding a known volume of a silver nitrate solution. The mass of the precipitate is then measured, and the mass of the chloride ions in the original solution is calculated. The experiment is performed in a fume hood to avoid exposure to the toxic silver nitrate. The experiment is performed in a fume hood to avoid exposure to the toxic silver nitrate. The experiment is performed in a fume hood to avoid exposure to the toxic silver nitrate.

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Procedure:

The procedure involves the gravimetric analysis of a solution containing chloride ions. The principle of gravimetric analysis is based on the fact that the mass of a substance is proportional to the number of moles of the substance. In this experiment, the chloride ions in the solution are precipitated as silver chloride (AgCl) by adding a known volume of a silver nitrate solution. The mass of the precipitate is then measured, and the mass of the chloride ions in the original solution is calculated.

Data Analysis:

	Sample 1	Sample 2
Mass of sample	0.100 g	0.100 g
Mass of precipitate	0.150 g	0.150 g
Mass of chloride ion	0.050 g	0.050 g

The mass of the chloride ion in the sample is calculated as follows:

$$\text{Mass of chloride ion} = \text{Mass of precipitate} - \text{Mass of silver ion}$$

The mass of the silver ion in the sample is calculated as follows:

Analysis Of Chloride Experiment 7 Lab Report

Joacim Rocklöv



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