

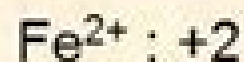
## Balancing redox equations

### Example

Balance the equation showing the oxidation of  $\text{Fe}^{2+}$  ions to  $\text{Fe}^{3+}$  ions by dichromate ions ( $\text{Cr}_2\text{O}_7^{2-}$ ) in an acidic medium



Step 1: Identify oxidising and reducing agents and write half reactions



$$2\text{Cr} = 12$$

$$\text{Cr} = +6$$



Iron loses one electron, i.e. it is *oxidised*. It donates one electron to dichromate, i.e. it is the *reducing agent*.

Half equation: Oxidation reaction:  $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$

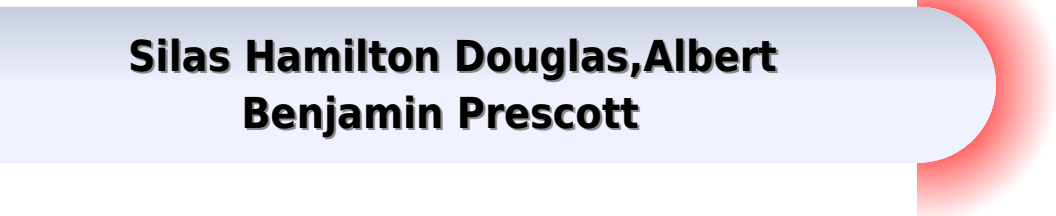


Chromium gains three electrons, i.e. it is *reduced*. It gains its electrons from iron, i.e. it is the *oxidising agent*.

Half equation: Reduction reaction:  $\text{Cr}_2^{6+} + 3\text{e}^- \rightarrow \text{Cr}^{3+}$

# Analyzing Oxidation Reduction Reactions Answers

**Silas Hamilton Douglas, Albert  
Benjamin Prescott**



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