

Lab 10: Calculations with a Chemical Structure

Objectives:

1. To derive an empirical formula from a chemical structure without performing stoichiometric calculations.
2. To calculate the number of moles of a solid from a weighing measurement in the lab.
3. To determine the molar mass of a solid.
4. To determine the stoichiometric amount of a solid reagent that would be produced.
5. To compare the theoretical amount to the actual amount of a substance formed and calculate the percent yield.

Reagents and Equipment:

Reagent of Sodium Chloride solution: 0.5 M and

Reagent of Calcium Chloride solution: 0.5 M and

Reagent of Silver Nitrate solution: 0.5 M and

Reagent of Barium Chloride solution: 0.5 M and

Reagent of Potassium Nitrate solution: 0.5 M and

Reagent of Sodium Chloride solution: 0.5 M and

Reagent of Calcium Chloride solution: 0.5 M and

Procedure:

Reagents and Calculations:

1. Calculations:

$$\text{moles of sodium chloride} = \frac{\text{mass of sodium chloride}}{\text{molar mass of sodium chloride}} = \frac{0.5 \text{ g}}{58.44 \text{ g/mol}} = 0.00855 \text{ mol}$$

$$\text{moles of calcium chloride} = \frac{\text{mass of calcium chloride}}{\text{molar mass of calcium chloride}} = \frac{0.5 \text{ g}}{147.01 \text{ g/mol}} = 0.00339 \text{ mol}$$

$$\text{moles of sodium chloride produced} = 0.00855 \text{ mol} + 0.00339 \text{ mol} = 0.01194 \text{ mol}$$

$$\text{moles of calcium chloride produced} = 0.00339 \text{ mol} + \frac{0.00855 \text{ mol}}{2} = 0.00767 \text{ mol}$$

2. Chemical Equations: $\text{Na}_2\text{CO}_3 + \text{CaCl}_2 \rightarrow \text{CaCO}_3 + 2\text{NaCl}$

3. Reaction Equations:

$$\text{Na}_2\text{CO}_3 + \text{CaCl}_2 \rightarrow \text{CaCO}_3 + 2\text{NaCl} \quad \text{This reaction is used to determine the amount of calcium chloride produced.}$$

$$\text{Na}_2\text{CO}_3 + \text{CaCl}_2 \rightarrow \text{CaCO}_3 + 2\text{NaCl}$$

$$\text{moles of Na}_2\text{CO}_3 \text{ reacted} = \frac{\text{mass of Na}_2\text{CO}_3}{\text{molar mass of Na}_2\text{CO}_3} = \frac{0.5 \text{ g}}{105.99 \text{ g/mol}} = 0.00472 \text{ mol}$$

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Ying-Ying Zheng



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