

Stoichiometry Worksheet #2 (mole-mass, mass-mole problems)



- a. If 15.0g of N_2O_4 was produced, how many moles of O_2 were required?

$$\frac{15.0\text{g N}_2\text{O}_4}{92.0\text{g N}_2\text{O}_4} \times \frac{1\text{ mol N}_2\text{O}_4}{1\text{ mol N}_2\text{O}_4} \times \frac{2\text{ mol O}_2}{1\text{ mol N}_2\text{O}_4} = 0.326\text{ mol O}_2$$

- b. If 4.0×10^{-3} moles of oxygen reacted, how many grams of N_2 were needed?

$$\frac{4.0 \times 10^{-3}\text{ mol O}_2}{2\text{ mol O}_2} \times \frac{1\text{ mol N}_2}{1\text{ mol N}_2} \times \frac{28.0\text{g N}_2}{1\text{ mol N}_2} = 5.6 \times 10^{-2}\text{ g N}_2$$



- a. What is the mass of potassium nitrate that is produced when 2.04 moles of potassium phosphate react?

$$\frac{2.04\text{ mol K}_3\text{PO}_4}{1\text{ mol K}_3\text{PO}_4} \times \frac{3\text{ mol KNO}_3}{1\text{ mol K}_3\text{PO}_4} \times \frac{101.1\text{ g KNO}_3}{1\text{ mol KNO}_3} = 619\text{g KNO}_3$$

- b. If 5.80g of aluminum phosphate are formed, how many moles of aluminum nitrate reacted?

$$\frac{5.80\text{g AlPO}_4}{122.0\text{g AlPO}_4} \times \frac{1\text{ mol AlPO}_4}{1\text{ mol AlPO}_4} \times \frac{1\text{ mol Al}(\text{NO}_3)_3}{1\text{ mol AlPO}_4} = 4.75 \times 10^{-2}\text{ mol Al}(\text{NO}_3)_3$$



- a. If you have 5.50 mol of CaC_2 , how much C_2H_2 do you get?

$$\frac{5.50\text{ mol CaC}_2}{1\text{ mol CaC}_2} \times \frac{1\text{ mol C}_2\text{H}_2}{1\text{ mol CaC}_2} \times \frac{26.0\text{g C}_2\text{H}_2}{1\text{ mol C}_2\text{H}_2} = 143\text{g C}_2\text{H}_2$$

- b. How many moles of water are needed when 65.0g of CaC_2 have reacted?

$$\frac{65.0\text{g CaC}_2}{64.1\text{g CaC}_2} \times \frac{1\text{ mol CaC}_2}{1\text{ mol CaC}_2} \times \frac{2\text{ mol H}_2\text{O}}{1\text{ mol CaC}_2} = 2.03\text{ mol H}_2\text{O}$$

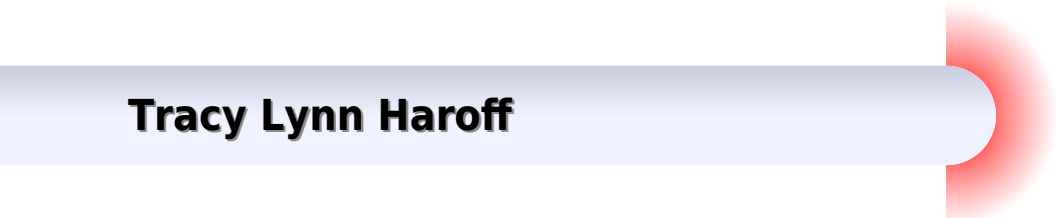
4. In photosynthesis, water reacts with carbon dioxide to give oxygen and glucose ($\text{C}_6\text{H}_{12}\text{O}_6$). Write and balance the chemical equation. How many moles of CO_2 are required to make 120.0g of glucose?



$$\frac{120.0\text{g C}_6\text{H}_{12}\text{O}_6}{180.0\text{g C}_6\text{H}_{12}\text{O}_6} \times \frac{1\text{ mol C}_6\text{H}_{12}\text{O}_6}{1\text{ mol C}_6\text{H}_{12}\text{O}_6} \times \frac{6\text{ mol CO}_2}{1\text{ mol C}_6\text{H}_{12}\text{O}_6} = 4.000\text{ mol CO}_2$$

Answers For The Chemistry Lab Of Stoichiometry

Tracy Lynn Haroff



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Optimizing STEM Education With Advanced ICTs and Simulations Levin, Ilya, Tsybulsky, Dina, 2017-06-05 The role of technology in educational settings has become increasingly prominent in recent years When utilized effectively these tools provide a higher quality of learning for students Optimizing STEM Education With Advanced ICTs and Simulations is an innovative reference source for the latest scholarly research on the integration of digital tools for enhanced STEM based learning environments Highlighting a range of pivotal topics such as mobile games virtual labs and participatory simulations this publication is ideally designed for educators professionals academics and students seeking material on emerging educational technologies

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chemistry Experiments are written to address students of various academic backgrounds and differing interests and abilities in chemistry While most experiments can be conducted in a single three hour period some have been designed to be completed over an extended time to illustrate that chemical systems do not work at an arbitrary schedule Suggestions are provided for combining experiments of shorter length and similar pedagogy Improving Student Comprehension of Stoichiometric Concepts Connie Lynn Bannick Kemner,2007 **Addison-Wesley Small-scale Chemistry** Dennis D. Staley,Edward L. Waterman,1995 Cooperative Chemistry Lab Manual Cooper,2005-02 The laboratory course described in the lab manual emphasizes experimental design data analysis and problem solving Inherent in the design is the emphasis on communication skills both written and oral Students work in groups on open ended projects in which they are given an initial scenario and then asked to investigate a problem There are no formalized instructions and students must plan and carry out their own investigations **Pp/Chemistry** BarCharts, Inc.,2008-06-18 **15 PGT Chemistry Test Papers EMRS** Mocktime Publication, EMRS Exam Teachers PGT Chemistry Test Papers 15 Practice Papers Tier 1 Eklayva Model Residential Schools as per Official Exam Pattern and Syllabus **SAT Subject Test: Chemistry with Online Tests** Joseph A. Mascetta,Mark Kernion,2018-09-01 Always study with the most up to date prep Look for SAT Subject Test Chemistry ISBN 9781506263120 on sale December 01 2020 Publisher s Note Products purchased from third party sellers are not guaranteed by the publisher for quality authenticity or access to any online entitles included with the product *Illustrated Guide to Home Chemistry Experiments* Robert Bruce Thompson,2008-04-29 For students DIY hobbyists and science buffs who can no longer get real chemistry sets this one of a kind guide explains how to set up and use a home chemistry lab with step by step instructions for conducting experiments in basic chemistry not just to make pretty colors and stinky smells but to learn how to do real lab work Purify alcohol by distillation Produce hydrogen and oxygen gas by electrolysis Smelt metallic copper from copper ore you make yourself Analyze the makeup of seawater bone and other common substances Synthesize oil of wintergreen from aspirin and rayon fiber from paper Perform forensics tests for fingerprints blood drugs and poisons and much more From the 1930s through the 1970s chemistry sets were among the most popular Christmas gifts selling in the millions But two decades ago real chemistry sets began to disappear as manufacturers and retailers became concerned about liability em The Illustrated Guide to Home Chemistry Experiments steps up to the plate with lessons on how to equip your home chemistry lab master laboratory skills and work safely in your lab The bulk of this book consists of 17 hands on chapters that include multiple laboratory sessions on the following topics Separating Mixtures Solubility and Solutions Colligative Properties of Solutions Introduction to Chemical Reactions Stoichiometry Reduction Oxidation Redox Reactions Acid Base Chemistry Chemical Kinetics Chemical Equilibrium and Le Chatelier s Principle Gas Chemistry Thermochemistry and Calorimetry Electrochemistry Photochemistry Colloids and Suspensions Qualitative Analysis Quantitative Analysis Synthesis of Useful Compounds Forensic Chemistry With plenty of full color illustrations and photos Illustrated Guide to

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