

OpenQuest 13

Orbital Configuration

Name _____

Date _____

Class _____

Information: Energy of Sublevels

Not all orbitals have the same energy of energy. For example, orbitals in the $3p$ sublevel have more energy than orbitals in the $3s$ sublevel. The following orbitals have sublevels from lowest to highest energy:



Finally, you have information for which orbitals are filled. Here's the orbital configuration for Boron:



Now the 4 orbitals appear in the set of orbitals. These orbitals have the same energy level. For example, the 4 orbitals have higher energy than the 3 orbitals, but the 4 orbitals have a higher energy than the 3 orbitals.

When electrons occupy orbitals, they try to have the lowest possible energy possible. (This is called the **lowest energy principle**.) An electron will occupy a 4 orbital before the 3 orbitals are filled, and an electron will occupy a 4 orbital before the 3 orbitals are filled. Additionally, an electron will occupy a 4 orbital. (This is called the **Pauli Exclusion Principle**.) When two electrons occupy the same orbital, they must spin in opposite directions—orbital electron configuration is opposite.

Critical Thinking Questions

- a. How many electrons would it take to fill the 4 orbitals? (The 4 orbitals are filled, all the orbitals are filled, and the orbitals are filled.)
b. How many electrons would it take to fill the 4 orbitals? (The 4 orbitals are filled, all the orbitals are filled, and the orbitals are filled.)
- a. How many electrons would it take to fill the 4 orbitals? (The 4 orbitals are filled, all the orbitals are filled, and the orbitals are filled.)
b. How many electrons would it take to fill the 4 orbitals? (The 4 orbitals are filled, all the orbitals are filled, and the orbitals are filled.)

Information: Hund's Rule

Electrons prefer to "split" or "spread". Electrons in the same orbital will have the same energy level. Electrons in the same orbital will have the same energy level. Electrons in the same orbital will have the same energy level.



Electrons in the same orbital will have the same energy level. Electrons in the same orbital will have the same energy level. Electrons in the same orbital will have the same energy level.



Hund's rule is to split the electrons into orbitals. Hund's rule is to split the electrons into orbitals. Hund's rule is to split the electrons into orbitals.

The following are the electron orbital diagrams for the elements, nitrogen and oxygen. Under the diagram's 4 orbitals are filled, and the 4 orbitals are filled.



Critical Thinking Questions

- a. How many "split" electrons are in a nitrogen atom?
- b. Why does nitrogen's electron configuration is $1s^2, 2s^2, 2p^3$? Why not $1s^2, 2s^2, 2p^4$?
- c. Why does oxygen's electron configuration is $1s^2, 2s^2, 2p^4$?
- d. Why does nitrogen's electron configuration is $1s^2, 2s^2, 2p^3$?
- e. Why does oxygen's electron configuration is $1s^2, 2s^2, 2p^4$?

Answers For Chemquest

M Tight



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