

Chemistry Practice Exam v2

Directions: It is important that you provide answers in your own words. Please focus only on information from the text/eBook to create your own solutions. Please do not use direct information from an outside source (especially copying and pasting from an "answer" website). Use of direct information from an outside source is against school policy. *All answers will be checked for plagiarism. Instances of plagiarism can result in probation or possible dismissal from the school.*

Please be sure to follow all guidelines (number of sentences/showing all calculations) and to provide the correct metric units of measure. All questions are 3.33 points each (.33 for unit)

1. A student collected the following data for a fixed volume of gas:

Temperature ($^{\circ}\text{C}$)	Pressure (mm of Hg)
10	726
20	750
40	800
70	880
100	960
150	1085

Fill in the missing data point. Show all calculations leading to an answer.

The General Law of Gases states that

$$P_1 \cdot V_1 / T_1 = P_2 \cdot V_2 / T_2.$$

As Volume is constant, we can write this as

$$P_1 / T_1 = P_2 / T_2.$$

P_1 and T_1 are the pressure and temperature at time 1.

P_2 and T_2 are the pressure and temperature at time 2.

You can take any pair of values as P_1 and T_1 and find $P_2 = P_1 T_2 / T_1$, once you sum 273.15 to each temperature to get kelvins.

$$P_1 = 726 \text{ mm Hg}$$

$$T_1 = 10 + 273.15 = 283.15 \text{ K};$$

$$T_2 = 150 + 273.15 = 423.15 \text{ K}.$$

$$\text{Then, } P_2 = 726 \cdot 423.15 / 283.15 = 1085.96 \text{ mm Hg}.$$

2. You are given a clear solution of KNO_3 . Using 3 – 4 sentences (in your own words) explain how you would determine if the solution is unsaturated, saturated or supersaturated.

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Jessica J Manson



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Table of Contents Answers Modeling Chemistry U7 Empfrm Practice V2 0

1. Understanding the eBook Answers Modeling Chemistry U7 Empfrm Practice V2 0
 - The Rise of Digital Reading Answers Modeling Chemistry U7 Empfrm Practice V2 0
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers Modeling Chemistry U7 Empfrm Practice V2 0
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers Modeling Chemistry U7 Empfrm Practice V2 0
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers Modeling Chemistry U7 Empfrm Practice V2 0
 - Personalized Recommendations
 - Answers Modeling Chemistry U7 Empfrm Practice V2 0 User Reviews and Ratings
 - Answers Modeling Chemistry U7 Empfrm Practice V2 0 and Bestseller Lists
5. Accessing Answers Modeling Chemistry U7 Empfrm Practice V2 0 Free and Paid eBooks
 - Answers Modeling Chemistry U7 Empfrm Practice V2 0 Public Domain eBooks
 - Answers Modeling Chemistry U7 Empfrm Practice V2 0 eBook Subscription Services
 - Answers Modeling Chemistry U7 Empfrm Practice V2 0 Budget-Friendly Options

6. Navigating Answers Modeling Chemistry U7 Empfrm Practice V2 0 eBook Formats
 - ePub, PDF, MOBI, and More
 - Answers Modeling Chemistry U7 Empfrm Practice V2 0 Compatibility with Devices
 - Answers Modeling Chemistry U7 Empfrm Practice V2 0 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers Modeling Chemistry U7 Empfrm Practice V2 0
 - Highlighting and Note-Taking Answers Modeling Chemistry U7 Empfrm Practice V2 0
 - Interactive Elements Answers Modeling Chemistry U7 Empfrm Practice V2 0
8. Staying Engaged with Answers Modeling Chemistry U7 Empfrm Practice V2 0
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers Modeling Chemistry U7 Empfrm Practice V2 0
9. Balancing eBooks and Physical Books Answers Modeling Chemistry U7 Empfrm Practice V2 0
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers Modeling Chemistry U7 Empfrm Practice V2 0
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answers Modeling Chemistry U7 Empfrm Practice V2 0
 - Setting Reading Goals Answers Modeling Chemistry U7 Empfrm Practice V2 0
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answers Modeling Chemistry U7 Empfrm Practice V2 0
 - Fact-Checking eBook Content of Answers Modeling Chemistry U7 Empfrm Practice V2 0
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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