

SOLUBILITY CURVE WORKSHEET KEY

Use your solubility curve graphs provided to answer the following questions.

1. What are the customary units of solubility on solubility curves? Degress Celsius and grams of solute/100g of water
2. Define solubility. A measure of how much solute can dissolve in a given amount of solvent.
3. According to the graph, the solubility of any substance changes as temperature changes.
4. List the substances whose solubility decreases as temperature increases. NH_3 and $\text{Ce}_2(\text{SO}_4)_3$
5. Which substance is least affected by temperature changes? NaCl
6. How many grams of ammonium chloride (NH_4Cl) at 50°C ? 50g
7. NaCl and KClO_3 have the same solubility at approximately 78°C .
8. Which compound is least soluble in water at 10°C ? KClO_3
9. How many grams of KNO_3 can be dissolved at 50°C ? 80g
10. Are the following solutions unsaturated, saturated, or supersaturated?
 - a. 45g of NaNO_3 in 100 g of water at 30°C . saturated
 - b. 60g of KClO_3 in 100 g of water at 90°C . supersaturated
11. How many grams of sodium chloride, NaCl are required to saturate 100 grams of water at 100°C ? 40g
12. How many grams of NaNO_3 are required to saturate 100 grams of water at 75°C ? 140g
13. How many grams of KI will saturate water at 20°C ? 33g
14. At what temperature would 25g of potassium chlorate (KClO_3) dissolve? 60°C
15. At what temperature would 60g of NH_4Cl dissolve? 70°C
16. 89 g NaNO_3 is prepared at 30°C .
 - a) Will all of the salt dissolve? No
 - b) What mass of NaNO_3 will dissolve at this temperature? 95g
17. If 50 grams of NH_4Cl is dissolved at 50°C , how many additional grams NH_4Cl would be needed to make the solution saturated at 80°C ? 15g
18. At 50°C , how many grams of KNO_3 will dissolve? 80g
19. At 70°C , how many grams of cerium (III) sulfate ($\text{Ce}_2(\text{SO}_4)_3$) dissolve? 5g
20. Determine if each of the following is unsaturated, saturated, or supersaturated.
 - a. 55g of NH_3 at 20°C . supersaturated
 - b. 10g of $\text{Ce}_2(\text{SO}_4)_3$ at 10°C . unsaturated
 - c. 110g of KNO_3 at 60°C . supersaturated
 - d. 65g of NH_4Cl at 80°C . saturated
 - e. 12g of NH_3 at 90°C . supersaturated
 - f. 78g of NaNO_3 at 10°C . saturated
 - g. 145g of NaNO_3 at 80°C . saturated
 - h. 35g of NaCl at 100°C . unsaturated

Answer Key To Solubility Curve 1

Brian Pressley

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Answer Key To Solubility Curve 1:

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Colligative Properties Electrochemistry Chemical Kinetics Nuclear Chemistry Surface Chemistry s block Elements p block Elements I p block Elements II Transition and Inner Transition Elements Coordination Compounds Extraction of Metals Qualitative Analysis Organic Chemistry Basics Hydrocarbons Alkyl Halides Alcohols and Ethers Aldehydes and Ketones Carboxylic Acids and their Derivatives Aliphatic Compounds Containing Nitrogen Benzene and Alkyl Benzene Aromatic Compounds Containing Nitrogen Aryl Halides and Phenols Aromatic Aldehydes Ketones and Acids Biomolecules and Chemistry in Everyday Life Environmental Chemistry JEE Advanced Solved Paper 2021 Chemistry, 2015-03-16 Chemistry for grades 9 to 12 is designed to aid in the review and practice of chemistry topics Chemistry covers topics such as metrics and measurements matter atomic structure bonds compounds chemical equations molarity and acids and bases The book includes realistic diagrams and engaging activities to support practice in all areas of chemistry The 100 Series science books span grades 5 to 12 The activities in each book reinforce essential science skill practice in the areas of life science physical science and earth science The books include engaging grade appropriate activities and clear thumbnail answer keys Each book has 128 pages and 100 pages or more of reproducible content to help students review and reinforce essential skills in individual science topics The series will be aligned to current science standards **Chemistry** Carson-Dellosa Publishing, 2015-03-16 Chemistry for grades 9 to 12 is designed to aid in the review and practice of chemistry topics Chemistry covers topics such as metrics and measurements matter atomic structure bonds compounds chemical equations molarity and acids and bases The book includes realistic diagrams and engaging activities to support practice in all areas of chemistry The 100 Series science books span grades 5 to 12 The activities in each book reinforce essential science skill practice in the areas of life science physical science and earth science The books include engaging grade appropriate activities and clear thumbnail answer keys Each book has 128 pages and 100 pages or more of reproducible content to help students review and reinforce essential skills in individual science topics The series will be aligned to current science standards **Research and Development Progress Report** United States. Office of Saline Water, 1969 Science Activity Book Chapterwise Class 12 Priti Singhal, 2024-11-17 This book is designed to ignite curiosity and foster a love for science in students from grades 1 to 12 With a diverse range of engaging activities this book aims to provide a hands on interactive approach to understanding fundamental scientific concepts tailored to the unique developmental stages across all grade levels Our primary goal is to make learning science enjoyable and enriching The book is filled with colourful illustrations real life examples and interactive exercises that help students understand and relate to the world around them Each chapter is carefully structured to build on prior knowledge ensuring a steady progression in learning as students advance through the grades **Quarterly Journal of the Chemical Society of London** Chemical Society (Great Britain), 1958-04 *Crystallization* Marcello Andreeta, 2012-09-19 Crystallization is one of the most ancient and interdisciplinary topics of research known to mankind Crystals can be organic or inorganic and may be produced from melts

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Atul,2010-09 **Chemical Engineering for Non-Chemical Engineers** Jack Hipple,2017-02-06 Outlines the concepts of chemical engineering so that non chemical engineers can interface with and understand basic chemical engineering concepts Overviews the difference between laboratory and industrial scale practice of chemistry consequences of mistakes and approaches needed to scale a lab reaction process to an operating scale Covers basics of chemical reaction engineering mass energy and fluid energy balances how economics are scaled and the nature of various types of flow sheets and how they are developed vs time of a project Details the basics of fluid flow and transport how fluid flow is characterized and explains the difference between positive displacement and centrifugal pumps along with their limitations and safety aspects of these differences Reviews the importance and approaches to controlling chemical processes and the safety aspects of controlling chemical processes Reviews the important chemical engineering design aspects of unit operations including distillation absorption and stripping adsorption evaporation and crystallization drying and solids handling polymer manufacture and the basics of tank and agitation system design **Mass Transfer Operations** Mr. Rohit Manglik,2024-07-24 Discusses theoretical and practical aspects of mass transfer including distillation absorption and drying for students of chemical engineering **Handbook of Industrial Crystallization** Allan Myerson,2002-01-08 Crystallization is an important

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