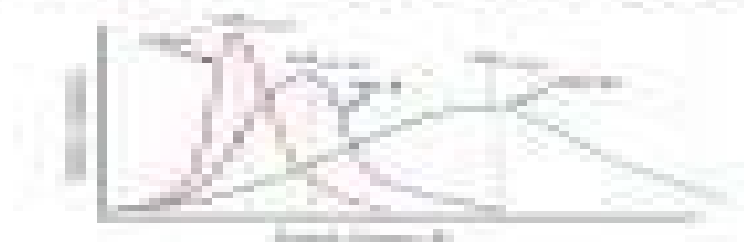


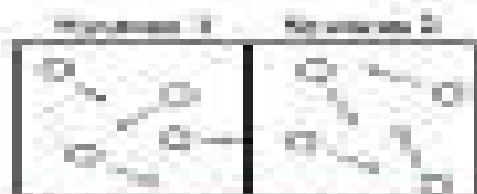
Unit 5: Thermochemistry**5.1: Heat Transfer and Thermal Equilibrium****5.2: Heat Capacity and Calorimetry****5.3: Entropy and Spontaneous Changes**

2. Use the Boltzmann probability distribution curves below to predict temperatures of the mixtures of particles.



Temperature: is a measure of the average kinetic energy of the particles in a system. The kinetic energy for particles moving with m in cm^2/s^2 is the mass of particles and momentum mass and change, indirectly connected to all of temperature increase. The Boltzmann probability curves show that the average velocity of each of the distributions is increasing (less to large) and the average temperature is increasing.

3. Sketch a curve to represent the relative distribution of the particles in the following two parts: samples of the same gas.



- a. In System 1, it is a higher or lower temperature than System 2? Justify your answer.
 System 1 is at a lower temperature than System 2. Temperature is a measure of the average KE of the particles in the system. In System 1, the particles are more closely packed, and all the intermolecular interactions within the system are stronger, resulting in a lower average KE. In System 2, the particles are more spread out, and the intermolecular interactions are weaker, resulting in a higher average KE.
- b. An initial distribution with kinetic energy for System 1. Justify your answer in terms of molecular collisions.
 All particles in System 1 are at a lower temperature than System 2. The particles in System 1 are more closely packed, and the intermolecular interactions are stronger, resulting in a lower average KE. In System 2, the particles are more spread out, and the intermolecular interactions are weaker, resulting in a higher average KE.

Answer Key Unit 6 Thermochemistry

N Noddings



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